

Name: _____

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Unit: Acids, Bases, and Solutions

Acid and Base Calculations – Wksh #4



1. Calculate the pH of the solutions, given the following $[H^+]$, and then identify the solution as acidic, basic, or neutral.

a) $[H^+] = 1.2 \times 10^{-2} \text{ M}$

b) $[H^+] = 5.8 \times 10^{-6} \text{ M}$

c) $[H^+] = 3.92 \times 10^{-13} \text{ M}$

d) $[H^+] = 4.52 \times 10^{-7} \text{ M}$

2. Calculate the $[H^+]$ in the following solutions, given their pH, and identify the solution as acidic, basic or neutral.

a) $\text{pH} = 2.3$

b) $\text{pH} = 11.7$

c) $\text{pH} = 6.8$

d) $\text{pH} = 3.31$

3. Given the following pOH's, calculate the pH for the solutions:

a) $\text{pOH} = 4.18$

b) $\text{pOH} = 12.35$

4. Given the following pH's, calculate the pOH for the solutions:

a) $\text{pH} = 6.2$

b) $\text{pH} = 9.55$

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H. Stephen Stoker



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Chemistry Lemay, Prentice-Hall Staff, 1996 **Students Study Guide** Fredeen, Pearson, 2003-07 **Rocky Mountain Druggist**, 1930 Druggists' Circular, 1909 **American Druggists' Circular and Chemical Gazette**, 1922 **The Pharmaceutical Era**, 1898 *Boston Journal of Chemistry and Pharmacy*, 1893 **Lab Exs In Prin Med Sci** Pearson Custom Publishing, 1990

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