

Exercise 2: Creating an Activity Series

Data Table 2. Redox Reactions of Copper, Lead, and Zinc.

| Redox Metal | Redox ID | Solution | Immediate Observations | 30-Minute Observations |
|-------------|----------|--------------|---|---|
| Cu | 0 | $Pb(NO_3)_2$ | No noticeable changes | None |
| | 0 | $Zn(NO_3)_2$ | No noticeable changes | None |
| Pb | 0 | $CuSO_4$ | Blue solution, a rust color appeared on the surface | Blue color faded and appeared to be rusty |
| | 0 | $ZnSO_4$ | None | None |
| Zn | 0 | $CuSO_4$ | Appeared to be dark, rusty, and fuzzy appearance, blue solution | Appeared to be rusted, was a dark colored almost black, less blue |
| | 0 | $Pb(NO_3)_2$ | Dark almost black, clear solution | Appeared to be fuzzy, dark, and grey colored, clear solution |

Data Table 3. Potential Redox Reactions and Chemical Equations.

| Metal and Aqueous Solution | Reaction Occurred? | Chemical Equation |
|----------------------------|--------------------|---|
| $Cu + Pb(NO_3)_2$ | No | No reaction |
| $Cu + Zn(NO_3)_2$ | No | No reaction |
| $Pb + CuSO_4$ | Yes | $Pb + CuSO_4 \rightarrow Cu + PbSO_4$ |
| $Pb + Zn(NO_3)_2$ | No | $Pb + Zn(NO_3)_2 \rightarrow Zn + Pb(NO_3)_2$ |
| $Zn + CuSO_4$ | Yes | $Zn + CuSO_4 \rightarrow Cu + ZnSO_4$ |
| $Zn + Pb(NO_3)_2$ | Yes | $Zn + Pb(NO_3)_2 \rightarrow Pb + Zn(NO_3)_2$ |

Activity Series Lab Report Answers

LP Steffe



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