

Temperature

°C, K

measure of average KE
of motion of particles

Heat

kJ, kcal (Cal) [1 kcal = 4.184 kJ]

measure of total energy transferred from
an object of high E → low E.

Note: A change in T is accompanied by a transfer of heat energy.

Specific heat (c or c_p):

(specific for each substance)

amount of energy required to raise the temperature of 1 g of that substance by 1°C.

change in heat = specific heat x mass x change in T

$$q = c_p \times m \times \Delta T$$

Units: heat absorbed or released, in J;

ΔT in °C or K; m in g; c in J/g°C

Molar heat capacity = heat capacity in J/mol°C instead of J/g°C

Don't forget: the size of 1°C = 1 K, so K can be used in this calculation in place of °C.

Examples

- How much heat energy is released to your body when a cup of hot tea containing 200. g of water is cooled from 65.0°C to body temperature, 37.0°C? (c_p of H₂O = 4.18 J/g°C)
- A 4.50-g nugget of pure gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25.0°C? (c_p of gold = 0.129 J/g°C)
- A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance?

Chapter 6 Thermochemistry Test

George M. Bodner, Harry L. Pardue

A red circular graphic with a gradient, appearing as a semi-circle on the right side of the slide, partially overlapping the grey bar.

Chapter 6 Thermochemistry Test:

Unveiling the Power of Verbal Beauty: An Mental Sojourn through **Chapter 6 Thermochemistry Test**

In a world inundated with screens and the cacophony of instant communication, the profound energy and mental resonance of verbal artistry usually disappear into obscurity, eclipsed by the constant barrage of sound and distractions. Yet, nestled within the musical pages of **Chapter 6 Thermochemistry Test**, a interesting function of fictional beauty that pulses with organic feelings, lies an remarkable journey waiting to be embarked upon. Written by a virtuoso wordsmith, that mesmerizing opus manuals viewers on a mental odyssey, gently revealing the latent potential and profound impact stuck within the elaborate web of language. Within the heart-wrenching expanse with this evocative analysis, we can embark upon an introspective exploration of the book is central subjects, dissect its fascinating publishing model, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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