

Stoichiometric Calculations

Mole to Mole Calculations

1. The octane present in gasoline burns according to the following equation.



- (a) How many moles of O_2 are needed to react fully with 7 moles of octane (C_8H_{18})?
- (b) How many moles of CO_2 can form from 5 mole of octane?
- (c) If the reaction is to be used to synthesize 8 moles of CO_2 , how many moles of oxygen are needed?
- (d) How many moles of carbon dioxide are formed if 17 moles of oxygen are reacted?
2. Butane combines with oxygen in a combustion reaction to produce carbon dioxide and water.
- (a) Write the balanced equation for this reaction.
- (b) If 37 moles of water are produced, how many moles of butane were reacted?
- (c) 29 moles of oxygen are reacted with the butane, how many moles of carbon dioxide are produced?
- (d) How many moles of butane are required to react with 17.5 moles of oxygen?
3. Consider the following reaction:



- (a) How many moles of NO are produced by the reaction of 4.0 moles of copper with HNO_3 ?
- (b) How many moles of HNO_3 are required to react completely with 5.0 mol of copper?
- (c) What mass of NO is produced when 6.35 g of copper is reacted with HNO_3 ?

Chapter 122 Stoichiometric Calculations

Pauline M. Doran



Chapter 122 Stoichiometric Calculations:

STOICHIOMETRY AND PROCESS CALCULATIONS, SECOND EDITION NARAYANAN, K. V., LAKSHMIKUTTY,

B., 2016-12-01 Designed as a textbook for the undergraduate students of chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering and safety engineering the chief objective of the book is to prepare students to make analysis of chemical processes through calculations and to develop systematic problem solving skills in them The text presents the fundamentals of chemical engineering operations and processes in a simple style that helps the students to gain a thorough understanding of chemical process calculations The book deals with the principles of stoichiometry to formulate and solve material and energy balance problems in processes with and without chemical reactions With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations The book is supplemented with Solutions Manual for instructors containing detailed solutions of all chapter end unsolved problems NEW TO THE SECOND EDITION Incorporates a new chapter on Bypass Recycle and Purge Operations Comprises updations in some sections and presents new sections on Future Avenues and Opportunities in Chemical Engineering Processes in Biological and Energy Systems Contains several new worked out examples in the chapter on Material Balance with Chemical Reaction Includes GATE questions with answers up to the year 2016 in Objective type questions KEY FEATURES SI units are used throughout the book All basic chemical engineering operations and processes are introduced and different types of problems are illustrated with worked out examples Stoichiometric principles are extended to solve problems related to bioprocessing environmental engineering etc Exercise problems more than 810 are organised according to the difficulty level and all are provided with answers **Industrial Stoichiometry** Warren Kendall Lewis, Arthur Hitchcock Radasch, 1926 Fuels and combustion Gas producers Sulfur compounds Metallurgy Crystallization

Chemistry & Chemical Reactivity John C. Kotz, Paul Treichel, Patrick A. Harman, 2003 Provides a broad overview of the principles of chemistry the reactivity of chemical elements and their compounds and the applications of chemistry Conveys a sense of chemistry as a field that not only has a lively history but also one that is currently dynamic with important new developments on the horizon *Schaum's Outline of Beginning Chemistry, Third Edition* David Goldberg, 2009-09-14 Tough Test Questions Missed Lectures Not Enough Time Fortunately for you there s Schaum s Outlines More than 40 million students have trusted Schaum s to help them succeed in the classroom and on exams Schaum s is the key to faster learning and higher grades in every subject Each Outline presents all the essential course information in an easy to follow topic by topic format You also get hundreds of examples solved problems and practice exercises to test your skills This Schaum s Outline gives you Practice problems with full explanations that reinforce knowledge Coverage of the most up to date

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Davies, 2013-12-03 All general chemistry students face similar challenges but they use their textbook differently to meet
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strictly as per NCERT with deeper understanding of concepts This book covers wider application of concepts with focus on JEE Mains The book contains about 300 handpicked JEE Mains questions for MCQ practice on covering each topic separately The book has hundreds of topic wise illustrations and solved examples The book has hundreds of topic and subtopic wise practice questions The book has actual class like presentation of concepts The book is aimed at self paced learning

Bioprocess Engineering Principles Pauline M. Doran,2013 The emergence and refinement of techniques in molecular biology has changed our perceptions of medicine agriculture and environmental management This textbook presents the principles of bioprocess engineering in a way that is accessible to biological scientists

Stoichiometry: Atomic Weights, Molecular Formulas, Microcosmic Magnitudes Leonard Kollender Nash,1966

Investigating Chemistry Matthew Johll,2008-12-22 In its new second edition *Investigating Chemistry A Forensic Science Perspective* remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package

Chemistry in Quantitative Language Christopher O. Oriakhi,2009 Problem solving is one of the most challenging aspects students encounter in general chemistry courses leading to frustration and failure Consequently many students become less motivated to take additional chemistry courses after their first year This book deals with calculations in general chemistry and its primary goal is to prevent frustration by providing students with innovative intuitive and systematic strategies to problem solving in chemistry The material addresses this issue by providing several sample problems with carefully explained step by step solutions for each concept Key concepts basic theories and equations are provided and worked examples are selected to reflect possible ways problems could be presented to students

Analytical Chemistry Gary D. Christian,1994-01-06 Extensively revised and updated this edition is concerned primarily with quantitative analysis techniques Describes how to design an analytical method how to obtain a laboratory sample that is representative of the whole and to prepare it for analysis what measurement tools are available automated analyses and the statistical significance of the analysis New and expanded topics include heterogeneous equilibria diode array spectrometers fiber optic sensors and solid phase extraction

Chemistry Steven S. Zumdahl,1997

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 5 $f(x)$ expand vertically by a factor of translate down units
 6 $f(x)$ compress

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 a $f(4)$ b 0 for each question decide if

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 1 find w $7.5h = n^2$

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