

Acids & Bases

Arrhenius Acids & Bases

- Arrhenius acid: produces H^+
- Arrhenius base: produces OH^-
- Only applies to aqueous solutions & not all bases have OH^-

Bronsted Lowry

- Acids = proton donor
- Bases = proton acceptor
- Works in non aqueous

Protons

- proton= H^+ = H_3O^+

Conjugate Acids & Bases

- H^+ transferred to OH^-

K_a and K_b

- Acid dissociation constant
- Base dissociation constant
- $K_a = \frac{[H_3O^+][A^-]}{[HA]}$
- $K_b = \frac{[BH][OH^-]}{[B]}$

Acid Strengths

- Strong acid $\rightarrow K_a > 1$
 - Assume 100% dissociation
 - Yields weak conjugate base
 - $HCl, HBr, H_2SO_4, HClO_4, HI, HNO_3$
- Weak acid $\rightarrow K_a < 1$
 - Yields strong conjugate base
- When calculating
 - Strong & strong \rightarrow add concentrations & use
 - Strong & weak \rightarrow use strong
 - Weak & weak \rightarrow compare K_a & use the stronger one
- Polyprotic acid
 - Multiple protons

Water

- Acid & base
- Amphoteric \rightarrow behave as either
- Autoionization
 - $2 H_2O \rightleftharpoons H_3O^+ + OH^-$
- $K_w = 1.0 \times 10^{-14}$ at 25 degrees celsius
 - Neutral water $[H_3O^+] = [OH^-] = 1.0 \times 10^{-7}$

pH & pOH

- $pH = -\log[H_3O^+]$
- $pOH = -\log[OH^-]$
- $pOH + pH = 14$
- $[H_3O^+] = 10^{-pH}$

Chapter 15 Review Acids Bases

Peipei Pang



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Table of Contents Chapter 15 Review Acids Bases

1. Understanding the eBook Chapter 15 Review Acids Bases
 - The Rise of Digital Reading Chapter 15 Review Acids Bases
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 15 Review Acids Bases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 15 Review Acids Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 15 Review Acids Bases
 - Personalized Recommendations
 - Chapter 15 Review Acids Bases User Reviews and Ratings
 - Chapter 15 Review Acids Bases and Bestseller Lists
5. Accessing Chapter 15 Review Acids Bases Free and Paid eBooks
 - Chapter 15 Review Acids Bases Public Domain eBooks
 - Chapter 15 Review Acids Bases eBook Subscription Services

- Chapter 15 Review Acids Bases Budget-Friendly Options
- 6. Navigating Chapter 15 Review Acids Bases eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 15 Review Acids Bases Compatibility with Devices
 - Chapter 15 Review Acids Bases Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 15 Review Acids Bases
 - Highlighting and Note-Taking Chapter 15 Review Acids Bases
 - Interactive Elements Chapter 15 Review Acids Bases
- 8. Staying Engaged with Chapter 15 Review Acids Bases
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 15 Review Acids Bases
- 9. Balancing eBooks and Physical Books Chapter 15 Review Acids Bases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 15 Review Acids Bases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 15 Review Acids Bases
 - Setting Reading Goals Chapter 15 Review Acids Bases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 15 Review Acids Bases
 - Fact-Checking eBook Content of Chapter 15 Review Acids Bases
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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