

# Reaction Energy

## Section 1

Answer **ALL** of the following questions in the space provided.

- The following reaction is exothermic. What is the value of  $\Delta H_{rxn}$ ?
- The following reaction is exothermic. What is the value of  $\Delta H_{rxn}$ ?



- Is energy being absorbed or is energy being released? Liquid  $\text{H}_2\text{O}$  temperature?
- Is the value of the equilibrium constant for the reaction a negative or a positive number?

Answer **ALL** of the following questions in the box to the left. Show all your work in the space provided.

- 10.000 g of water at  $20.0^\circ\text{C}$  absorbs a certain amount of energy when it is heated to  $100.0^\circ\text{C}$ .

11. How much heat energy is needed to heat 10.000 g of water from  $20.0^\circ\text{C}$  to  $100.0^\circ\text{C}$ ?

12. The molar heat of vaporization of water is  $40.7 \text{ kJ/mol}$ . How much heat energy is needed to vaporize 10.000 g of water at  $100.0^\circ\text{C}$ ?

# Chapter 16 Review Reaction Energy Section 1 Short Answer

**Lei Huang**



## Chapter 16 Review Reaction Energy Section 1 Short Answer:

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