

Table 8.3 Partial Orbital Diagrams and Electron Configurations* for the Elements in Period 3

Atomic Number/ Element	Partial Orbital Diagram (3s and 3p Sublevels Only)		Full Electron Configuration	Condensed Electron Configuration
11/Na	3s ↑	3p □ □ □	$[1s^2 2s^2 2p^6] 3s^1$	[Ne] $3s^1$
12/Mg	↑↓	□ □ □	$[1s^2 2s^2 2p^6] 3s^2$	[Ne] $3s^2$
13/Al	↑↓	↑ □ □	$[1s^2 2s^2 2p^6] 3s^2 3p^1$	[Ne] $3s^2 3p^1$
14/Si	↑↓	↑ ↑ □	$[1s^2 2s^2 2p^6] 3s^2 3p^2$	[Ne] $3s^2 3p^2$
15/P	↑↓	↑ ↑ ↑	$[1s^2 2s^2 2p^6] 3s^2 3p^3$	[Ne] $3s^2 3p^3$
16/S	↑↓	↑↓ ↑ ↑	$[1s^2 2s^2 2p^6] 3s^2 3p^4$	[Ne] $3s^2 3p^4$
17/Cl	↑↓	↑↓ ↑↓ ↑	$[1s^2 2s^2 2p^6] 3s^2 3p^5$	[Ne] $3s^2 3p^5$
18/Ar	↑↓	↑↓ ↑↓ ↑↓	$[1s^2 2s^2 2p^6] 3s^2 3p^6$	[Ne] $3s^2 3p^6$

Hund's rule

Chapter 8 Electron Configuration And General Chemistry

L Cohen



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Table of Contents Chapter 8 Electron Configuration And General Chemistry

1. Understanding the eBook Chapter 8 Electron Configuration And General Chemistry
 - The Rise of Digital Reading Chapter 8 Electron Configuration And General Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Electron Configuration And General Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Electron Configuration And General Chemistry
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chapter 8 Electron Configuration And General Chemistry
 - Personalized Recommendations
 - Chapter 8 Electron Configuration And General Chemistry User Reviews and Ratings
 - Chapter 8 Electron Configuration And General Chemistry and Bestseller Lists
5. Accessing Chapter 8 Electron Configuration And General Chemistry Free and Paid eBooks
 - Chapter 8 Electron Configuration And General Chemistry Public Domain eBooks
 - Chapter 8 Electron Configuration And General Chemistry eBook Subscription Services
 - Chapter 8 Electron Configuration And General Chemistry Budget-Friendly Options
6. Navigating Chapter 8 Electron Configuration And General Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Electron Configuration And General Chemistry Compatibility with Devices
 - Chapter 8 Electron Configuration And General Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Electron Configuration And General Chemistry
 - Highlighting and Note-Taking Chapter 8 Electron Configuration And General Chemistry
 - Interactive Elements Chapter 8 Electron Configuration And General Chemistry
8. Staying Engaged with Chapter 8 Electron Configuration And General Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Electron Configuration And General Chemistry
9. Balancing eBooks and Physical Books Chapter 8 Electron Configuration And General Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Electron Configuration And General Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Electron Configuration And General Chemistry
 - Setting Reading Goals Chapter 8 Electron Configuration And General Chemistry
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter 8 Electron Configuration And General Chemistry
 - Fact-Checking eBook Content of Chapter 8 Electron Configuration And General Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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