

Chapter 9: Chemical Reactions

- _____ 1. Which of the following could be an indication of a chemical reaction of a compound or element?
- It is a solid at room temperature.
 - It melts at 631^oC.
 - It expands upon freezing.
 - It burns in an atmosphere of chlorine.
- _____ 2. In this unbalanced chemical reaction, which components are the reactants?
- $$\text{C}_7\text{H}_{16} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$
- C_7H_{16} and O_2
 - C_7H_{16} and CO_2
 - C_7H_{16} and H_2O
 - CO_2 and H_2O
- _____ 3. When $\text{Ca}(\text{OH})_2$ reacts with H_2SO_4 , what are the products.
- $\text{Ca}(\text{OH})_2$ and H_2SO_4
 - CaH_2 and H_2SO_4
 - CaSO_4 and H_2O
 - CaO and H_2SO_4
- Hint: This is a double-replacement reaction, positive ions switch places.
- _____ 4. What are the products of the reaction between barium hydroxide and hydrochloric acid?
- Barium chloride and water
 - Barium hydroxide and water
 - Barium chloride and barium hydride
 - Chloric acid and barium hydroxide
- Hint: Write out the skeleton equation. This is a double-replacement reaction.
- _____ 5. This symbol separates the reactants from products and means "to yield".
- +
 - (s)
 -
 - (g)
- _____ 6. In a chemical equation, the number written in front of the chemical formula is the
- Reactant
 - Subscript
 - Equation
 - Coefficient
- _____ 7. When there are exactly the same number and type of atoms on both sides of a chemical equation, the equation is _____.
- Balanced
 - Ionic
 - Synthetic
 - Decomposed
- _____ 8. The chemical equation $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ is classified as a
- Synthesis reaction
 - Decomposition reaction
 - Single-replacement reaction
 - Double-replacement reaction
- _____ 9. In the combustion reaction of propane (C_3H_8) with oxygen gas, what are the products?
- C_3H_8 and O_2
 - C_3H_8 and CO_2
 - O_2 and H_2O
 - H_2O and CO_2
- _____ 10. The major physical indicator(s) of a combustion reaction is

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