

**CHAPTER 9 REVIEW***Stoichiometry***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- \_\_\_\_\_ The coefficients in a chemical equation represent the
    - masses in grams of all reactants and products.
    - relative number of moles of reactants and products.
    - number of atoms of each element in each compound in a reaction.
    - number of valence electrons involved in a reaction.
  - \_\_\_\_\_ Which of the following would not be studied within the topic of stoichiometry?
    - the mole ratio of Al to Cl in the compound aluminum chloride
    - the mass of carbon produced when a known mass of sucrose decomposes
    - the number of moles of hydrogen that will react with a known quantity of oxygen
    - the amount of energy required to break the ionic bonds in  $\text{CaF}_2$
  - \_\_\_\_\_ A balanced chemical equation allows you to determine the
    - mole ratio of any two substances in the reaction.
    - energy released in the reaction.
    - electron configuration of all elements in the reaction.
    - reaction mechanism involved in the reaction.
  - \_\_\_\_\_ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
    - reaction sequence.
    - bond energy.
    - mole ratio.
    - element proportion.
5. Given the reaction represented by the following unbalanced equation:  $\text{N}_2\text{O}(g) + \text{O}_2(g) \rightarrow \text{NO}_2(g)$
- Balance the equation.  
\_\_\_\_\_
  - What is the mole ratio of  $\text{NO}_2$  to  $\text{O}_2$ ?  
\_\_\_\_\_
  - If 20.0 mol of  $\text{NO}_2$  form, how many moles of  $\text{O}_2$  must have been consumed?  
\_\_\_\_\_
  - Twice as many moles of  $\text{NO}_2$  form as moles of  $\text{N}_2\text{O}$  are consumed. True or False?  
\_\_\_\_\_
  - Twice as many grams of  $\text{NO}_2$  form as grams of  $\text{N}_2\text{O}$  are consumed. True or False?  
\_\_\_\_\_

# Chapter 9 Review Stoichiometry Section 2 Work

**PT Brinkman**



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