

CHAPTER 1: STOICHIOMETRY

- One atom of sodium is
 - 23g
 - $\frac{1}{23}$ moles
 - Positively charged
 - $\frac{1}{23}$ moles
- Diamerck brown is a dye. Its molar mass is 228.3 g/mol. When the dye was analyzed by a scientist, it was found that it contains 50.50% nitrogen. How many nitrogen atoms are there in each Diamerck brown molecule?
 - 0
 - 5
 - 4
 - 3
- The mass of 2 moles of sodium hydroxide will be
 - 2g
 - 20g
 - 40g
 - 80g
- Consider the following reaction:

$$2Al_{(s)} + 3Cl_{2(g)} \rightarrow 2AlCl_{3(s)}$$

$2 \times 27 = 54$
 $3 \times 71 = 213$
 $54 + 213 = 267$
 $267 \times 2 = 534$

The following figure represents a mixture of $Al_{(s)}$ and $Cl_{2(g)}$.



Consider that each symbol is representing one mole of the substance. Which of the reactants will be consumed earlier and how many moles of the product will be formed?

- Al and 4 mol of $AlCl_3$
- Cl_2 and 4 mol of $AlCl_3$
- Cl_2 and 6 mol of $AlCl_3$
- both reactants will be consumed at the same time and 4 mol of $AlCl_3$

100 moles of H_2O contain
 100 moles of bonds
 100 moles of electrons

- 30 atoms
 - 40 moles of hydrogen bonds
- Ozone (O_3) filters the cosmic rays of sunlight. How many oxygen atoms are there in 0.2 mole of ozone (O_3 , molar mass 48.0 g/mol)?
 - 6.02×10^{23}
 - 6.02×10^{22}
 - 3.01×10^{23}
 - 6×10^{23}
 - On heating, the hydrated salt $CaSO_4 \cdot xH_2O$ loses its water of crystallization. In an experiment it was found that when 0.2 mol of the hydrated salt was heated 0.10 mol of water was lost. What is the molecular formula of the hydrated compound?
 - $CaSO_4$
 - $CaSO_4 \cdot 9H_2O$
 - $CaSO_4 \cdot H_2O$
 - $CaSO_4 \cdot 2H_2O$
 - Zirconium chloride reacts with magnesium to produce Zirconium

$$ZrCl_4 + 2Mg \rightarrow Zr + 2MgCl_2$$

3 mole 2 mole

How many moles of Zr would be produced if 0.2 mol of Mg metal is used in the reaction?

 - 0.1
 - 0.01
 - 0.20
 - 0.05
 - Tobacco contains a poisonous alkaloid called Nicotine. The molecular formula of nicotine is $C_{10}H_{14}N_2$ (molar mass = 162.23 g). The mass percentage of nitrogen in nicotine, is?
 - 4.32%
 - 8.23%
 - 12.4%
 - 17.3%
 - How much calcium is present in 10g of $CaCO_3$?
 - 1g
 - 2g
 - 3g
 - 4g
 - How many carbonate ions are there in 50g of $CaCO_3$?
 - 30 atoms
 - 40 moles of hydrogen bonds

$N_2 = 28g$
 1 mol

$$\frac{162.23}{28} \times 100 = 17.3\%$$

$0.2 \text{ mol } Mg / 2 \text{ mol } Mg \times 2 \text{ mol } Zr = 0.2 \text{ mol } Zr$

- $50 = 5.022 \times 10^{23}$
 - $100 = 5.022 \times 10^{23}$
 - $50 = 5.022 \times 10^{23}$
 - $50 = 5.022 \times 10^{23}$
- Stoichiometric calculations of when:
 - The reaction is irreversible
 - There is no side reaction
 - At least one of the reactants is completely consumed
 - All of the above
 - How many molecules of SO_2 form the decomposition of potassium chlorate by reaction?
 - $3 \times 6.022 \times 10^{23}$
 - 3×10^{23}
 - $3 \times 10^{23} \times 6.022 \times 10^{23}$
 - 6
 - Which of the following substances?
 - 4g helium
 - 6.022×10^{23} mol
 - 40g calcium
 - 6.022×10^{23}
 - Urea ($NH_2)_2CO$ and urea are present nitrogen in the compound?
 - 14g
 - 50g
 - 28g
 - 42g
 - 8g of methane
 - 4 hydrogens
 - 0.5 mol of carbon
 - 6.022×10^{23}
 - 4 \times 6.02
 - A beaker has 100 moles of oxygen
 - 10
 - 18



Chapter 9 Stoichiometry Multiple Choice Answers

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Chapter 9 Stoichiometry Multiple Choice Answers:

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