

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**2. PLAN**

- *What step is needed to calculate the concentration of the diluted solution?*

Apply the principle that $\text{volume}_1 \times \text{molarity}_1 = \text{volume}_2 \times \text{molarity}_2$.

$$M_1 V_1 = M_2 V_2 \xrightarrow[\text{substitute each of the known quantities for its symbol and calculate}]{\begin{array}{l} \text{rearrange the equation} \\ M_1 V_1 = M_2 V_2 \\ \text{algebraically to solve for } M_2 \end{array}} M_2 = \frac{M_1 V_1}{V_2}$$

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}}$$

3. COMPUTE

Note: Even though molarity is moles per liter, you can use volumes in milliliters along with molarity whenever the units cancel.

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}} = 0.948 \text{ M}$$

4. EVALUATE

- *Are the units correct?*
- *Is the number of significant figures correct?*
- *Is the answer reasonable?*

Yes; molarity (mol/L) was required.
 Yes; the number of significant figures is correct because the data had a minimum of three significant figures.
 Yes; the computation is the same as $4.74/5$, which is a little less than 1.

PRACTICE

1. Complete the table below by calculating the missing value in each row.

	Molarity of stock solution	Volume of stock solution	Molarity of dilute solution	Volume of dilute solution	
a.	0.500 M KBr	20.00 mL	7 M KBr	100.00 mL	ans: 0.100 M
b.	1.00 M LiOH	? mL	0.075 M LiOH	500.00 mL	ans: 38 mL
c.	7 M HI	5.00 mL	0.0493 M HI	100.00 mL	ans: 0.986 M
d.	12.0 M HCl	0.250 L	1.8 M HCl	? L	ans: 1.7 L
e.	7.44 M NH ₃	? mL	0.093 M NH ₃	4.00 L	ans: 50. mL

Chemfile Mini Guide To Problem Solving Answer Key

Rachel Turoscy



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Table of Contents Chemfile Mini Guide To Problem Solving Answer Key

1. Understanding the eBook Chemfile Mini Guide To Problem Solving Answer Key
 - The Rise of Digital Reading Chemfile Mini Guide To Problem Solving Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemfile Mini Guide To Problem Solving Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemfile Mini Guide To Problem Solving Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfile Mini Guide To Problem Solving Answer Key
 - Personalized Recommendations

- Chemfile Mini Guide To Problem Solving Answer Key User Reviews and Ratings
 - Chemfile Mini Guide To Problem Solving Answer Key and Bestseller Lists
5. Accessing Chemfile Mini Guide To Problem Solving Answer Key Free and Paid eBooks
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 - Chemfile Mini Guide To Problem Solving Answer Key eBook Subscription Services
 - Chemfile Mini Guide To Problem Solving Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chemfile Mini Guide To Problem Solving Answer Key Compatibility with Devices
 - Chemfile Mini Guide To Problem Solving Answer Key Enhanced eBook Features
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 - Highlighting and Note-Taking Chemfile Mini Guide To Problem Solving Answer Key
 - Interactive Elements Chemfile Mini Guide To Problem Solving Answer Key
 8. Staying Engaged with Chemfile Mini Guide To Problem Solving Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemfile Mini Guide To Problem Solving Answer Key
 9. Balancing eBooks and Physical Books Chemfile Mini Guide To Problem Solving Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemfile Mini Guide To Problem Solving Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemfile Mini Guide To Problem Solving Answer Key
 - Setting Reading Goals Chemfile Mini Guide To Problem Solving Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemfile Mini Guide To Problem Solving Answer Key
 - Fact-Checking eBook Content of Chemfile Mini Guide To Problem Solving Answer Key

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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