

## Chapter Test II

### A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

#### Column I

- \_\_\_\_\_ 1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
- \_\_\_\_\_ 2. The number of protons in the nucleus of an atom of an element.
- \_\_\_\_\_ 3. The number of electrons in a neutral atom.
- \_\_\_\_\_ 4. The ratio of the actual weight of an element to the standard weight of that element.
- \_\_\_\_\_ 5. The number of protons and neutrons in the nucleus of an atom.
- \_\_\_\_\_ 6. An element whose atoms have a mass number of 12.

#### Column II

- a. atomic weight
- b. limiting reagent
- c. molecular weight
- d. molar weight
- e. atomic number
- f. mass number

### B. Multiple Choice

Choose the correct answer for each question.

- \_\_\_\_\_ 1. The number of electrons in a neutral atom of an element is equal to the number of
  - a. protons in the nucleus
  - b. neutrons in the nucleus
  - c. atoms in the molecule
  - d. molecules in the sample
- \_\_\_\_\_ 2. Which of the following statements is true?
  - a. Molar weight is the sum of the atomic weights of the atoms in a molecule.
  - b.  $2 \text{ g oxygen} + 1 \text{ g hydrogen} = 3 \text{ g water}$ .
  - c.  $2 \text{ g oxygen} + 1 \text{ g hydrogen} = 3 \text{ g gas}$ .
  - d.  $2 \text{ g oxygen} + 1 \text{ g hydrogen} = 3 \text{ g water}$ .
  - e. None is true.
- \_\_\_\_\_ 3. The ratio of the actual weight of an element to the standard weight of that element is called
  - a. atomic weight
  - b. molecular weight
  - c. limiting reagent
  - d. molar weight
  - e. atomic number

# Chemistry Chapter 12 Stoichiometry Test

**David Baud**



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