

CHAPTER 3 CHAPTER ASSESSMENT

MATTER AND CHANGE

Reviewing Vocabulary

Match the definition in Column A with the term in Column B.

Column A

- _____ 1. The outermost electrons of an atom
- _____ 2. The center of an atom
- _____ 3. Hot, highly ionized, electrically conducting gas
- _____ 4. The smallest particle of an element that retains that element's characteristics
- _____ 5. The change from a gas to a liquid
- _____ 6. A homogeneous mixture
- _____ 7. The change of one or more substances into other substances
- _____ 8. The attractive force between two ions of opposite charge
- _____ 9. A tiny particle with a negative electrical charge
- _____ 10. The average of the mass numbers of an element's isotopes
- _____ 11. The spontaneous process through which unstable nuclei emit radiation

Column B

- a. atom
- b. atomic mass
- c. chemical reaction
- d. condensation
- e. electron
- f. ionic bond
- g. solution
- h. nucleus
- i. plasma
- j. valence electrons
- k. radioactive decay

In the space at the left, write the word or phrase in parentheses that makes the statement correct.

- _____ 12. The combined number of protons and neutrons is the (atomic number, mass number).
- _____ 13. An (energy level, isotope) represents the area in an atom where an electron is most likely to be found.
- _____ 14. (An ionic bond, A covalent bond) is the attraction of two atoms for a shared pair of electrons that holds the atoms together.
- _____ 15. An atom that gains or loses an electron is a charged particle called (an ion, a molecule).
- _____ 16. (An element, A compound) is a substance that cannot be broken down into simpler substances.
- _____ 17. A tiny particle that has mass and a positive electrical charge is a (neutron, proton).
- _____ 18. (Acids, Bases) are solutions characterized by the formation of hydroxide ions (OH^-).
- _____ 19. The process of changing from a liquid into a gas is called (evaporation, sublimation).

Chemistry Matter And Change Chapter 7

George Burton



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CHEMISTRY SILBERBERG,2003 **Chemistry in Action** George Mimms Rawlins,1961 Handbook for Waterworks Operator Certification Frank R. Spellman,2000-11-06 This three volume series is designed to prepare waterworks operators for certification and licensure exams Volume 1 is the only such volume based on the recently amended Safe Drinking Water Act and provides the tools to understand the microbiological and chemical hazards of water in light of the quality standards treatment plants must achieve Wit *Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed* Robert K. Wismer,1982 **Everyday Assessment in the Science Classroom** National Science Teachers Association,2003 The second in NSTA s Science Educator s Essay Collection Everyday Assessment is is designed to build confidence and enhance every teacher s ability to embed assessment into daily classwork The book s insights will help make assessment a dynamic classroom process of fine tuning how and what you teach **General, Organic, and Biological Chemistry** John R. Amend,Bradford P. Mundy,Melvin T. Arnold,1990 Very Good No Highlights or Markup all pages are intact *Chemical Ideas* George Burton,2000 This advanced chemistry text has been updated to match the specification for A Level Chemistry from September 2000 The problems have been revised and graded to allow more differentiation helping the teacher to teach students of a wide range of abilities The new editions of all the texts in this series should make it easier for teachers to match their teaching to the new modular specification There are new activities to cover ICT and key skills and end of unit tests to give students practice **Chemical Control of Gas Manufacture** ,1916 *Concepts of Materials Science* Adrian P. Sutton,2021-06-30 All technologies depend on the availability of suitable materials The progress of civilisation is often measured by the materials people have used from the stone age to the silicon age Engineers exploit the relationships between the structure properties and manufacturing methods of a material to optimise their design and production for particular applications Scientists seek to understand and predict those relationships This short book sets out fundamental concepts that underpin the science of materials and emphasizes their relevance to mainstream chemistry physics and biology These include the thermodynamic stability of materials in various environments quantum behaviour governing all matter and active matter Others include defects as the agents of change in crystalline materials materials at the nanoscale the emergence of new science at increasing length scales in materials and man made materials with properties determined by their structure rather than their chemistry The book provides a unique insight into the essence of materials science at a level suitable for pre university students and undergraduates of materials science It will also be suitable for graduates in other subjects contemplating postgraduate study in materials science Professional materials scientists will also find it stimulating and occasionally provocative **Pharmaceutical and Medical Chemistry** Samuel Philip Sadtler,Virgil Coblentz,Jeannot Hostmann,1927 *Understanding General Chemistry* Atef Korchef,2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the

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