

# Chemical Compounds and Bonding

## Solutions for Practice Problems

Student Textbook page 74

1. Determine  $\Delta EN$  for each bond shown. Indicate whether each bond is ionic or covalent.

(a) O–H	(b) C–H	(c) Mg–Cl	(d) B–F
(e) Cr–O	(f) C–N	(g) Na–I	(h) Na–Br

### What Is Required?

You have to determine the electronegativity difference between the bonded atoms listed and decide if it is an ionic or covalent bond.

### What Is Given?

The identity of each element in the chemical bond is given.

### Plan Your Strategy

Look up the electronegativity of each given element in Figure 3.6, student textbook page 71. For each bond combination, subtract the lower electronegativity value from the higher value to get the  $\Delta EN$  for the bond. From Figure 3.8, student textbook page 75, you can see that a  $\Delta EN$  value between 0.0 and 1.7 will likely be covalent, while a value between 1.7 and 3.3 will likely be ionic.

### Act on Your Strategy

- (a) O–H,  $\Delta EN = EN_O - EN_H = 3.44 - 2.20 = 1.24$ , covalent  
 (b) C–H,  $\Delta EN = EN_C - EN_H = 2.55 - 2.20 = 0.35$ , covalent  
 (c) Mg–Cl,  $\Delta EN = EN_{Cl} - EN_{Mg} = 3.16 - 1.31 = 1.85$ , ionic  
 (d) B–F,  $\Delta EN = EN_F - EN_B = 3.98 - 2.04 = 1.94$ , ionic  
 (e) Cr–O,  $\Delta EN = EN_O - EN_{Cr} = 3.44 - 1.66 = 1.78$ , ionic  
 (f) C–N,  $\Delta EN = EN_N - EN_C = 3.04 - 2.55 = 0.49$ , covalent  
 (g) Na–I,  $\Delta EN = EN_I - EN_{Na} = 2.66 - 0.93 = 1.73$ , ionic  
 (h) Na–Br,  $\Delta EN = EN_{Br} - EN_{Na} = 2.96 - 0.93 = 2.03$ , ionic

### Check Your Solution

An ionic bond will fall in the  $\Delta EN$  range of 1.7–3.3. A covalent bond will fall in the  $\Delta EN$  range of 0.0–1.7. Double check your electronegativity values from Figure 3.6. None of your answers should be negative values.

Student Textbook page 76

## 2. Problem

For each bond below, determine  $\Delta EN$ . Is the bond ionic or covalent?

(a) Ca–O	(b) K–Cl	(c) K–F
(d) Li–F	(e) Li–Br	(f) Ba–O

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**Jim Gaylor, May Look**



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