

**Standard-State Thermodynamic Values at 298.15 K:
Enthalpy of Formation (ΔH_f°), Free Energy of Formation (ΔG_f°), and
Absolute Entropy (S°)**

Substance	ΔH_f° (kJ/mol, ΔH_f°)	ΔG_f° (kJ/mol, ΔG_f°)	S° (J/mol·K, S°)
Aluminum			
Al(s)	0	0	28.3
Al ³⁺ (aq)	-1675.7	-1609.3	50.9
Al(OH) ₃ (s)	-1675.7	-1582.3	101.9
Al(OH) ₃ (aq)	-1675.7	-1582.3	118.67
Barium			
Ba(s)	0	0	62.5
Ba ²⁺ (aq)	-855.26	-762.97	99.6
BaCO ₃ (s)	-1362.3	-1194.3	101.42
BaCO ₃ (aq)	-1362.3	-1194.3	129.08
BaCl ₂ (s)	-858.59	-801.21	112.6
BaCl ₂ (aq)	-858.59	-801.21	112.6
Ba(OH) ₂ ·8H ₂ O(s)	-3302.67	-3026.94	213.94
Beryllium			
Be(s)	0	0	14.65
Be ²⁺ (aq)	-112.8	-112.8	12.9
BeCO ₃ (s)	-699.6	-699.6	129.7
BeCl ₂ (s)	-402.4	-402.4	102.04
Bismuth			
Bi(s)	0	0	96.24
Bi ₂ O ₃ (s)	-513.04	-492.7	112.9
BiCl ₃ (s)	-318.4	-318.4	117.6
Boron			
B(s)	0	0	5.9
B ₂ O ₃ (s)	-1273	-1273	201.81
Bromine			
Br ₂ (l)	0	0	152.23
Br ₂ (g)	30.91	3.11	245.46
Br ⁻ (aq)	-121.30	-112.54	82.4
HBr(g)	-36.40	-31.41	198.69
Carbon			
C(s, graphite)	0	0	57.07
C(s, diamond)	1.895	2.866	5.7
C(s, amorph)	0	0	5.7
CO(g)	-137.16	-137.16	197.67
CO ₂ (g)	-393.51	-394.36	213.74
Chlorine			
Cl ₂ (g)	0	0	223.08
Cl ⁻ (aq)	-167.16	-131.2	56.5
ClO ₂ (g)	102.82	52.01	220.18
ClO ₂ (aq)	102.82	52.01	61.9
ClO ₂ (l)	102.82	52.01	141.6

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S Marginson



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