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# Aiag Core Tools Manual

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 ec on the basis of this experiment what is the heat of solution per mole of  $KNO_3$  f w 101 11 u the specific heat of water is 4  
 184 j g k g soln 2 58 g 98 57 g 101 15 g  $\Delta t$  cal 20 4 ec 22 5 ec 2 1 ec

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web in a calorimeter the white solid  $CaSO_4$  forms and the temperature of the mixture increases to 28 1 c assuming that the

calorimeter absorbs only a negligible quantity of heat and that the specific heat capacity of the solution is  $4.18 \text{ J } ^\circ\text{C}^{-1} \text{ g}^{-1}$  and that the density of the final solution is  $1.0 \text{ g } \text{mL}^{-1}$  calculate the enthalpy change per mole

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