

Review for the Full Year General Chemistry  
American Chemical Society Exam

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

**THERMODYNAMICS**

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = 4.2 J g<sup>-1</sup> °C<sup>-1</sup>]

- (A) 0.11 J/g K (C) 17 J/g K  
(B) 0.22 J/g K (D) 34 J/g K

85. What is  $\Delta H$  in kJ for the reaction:

$[\Delta H_f^\circ \text{ (kJ/mol)}: \text{C}_2\text{H}_2(\text{g}) +227; \text{H}_2\text{O}(\text{g}) -242; \text{CO}_2(\text{g}) -393]$



- (A) -1830 kJ (B) -2284 kJ  
(C) -2510 kJ (D) -1605 kJ

86. Given these equations



calculate the heat of formation of SO<sub>3</sub>(g).

- (A) -500 kJ mol<sup>-1</sup> (C) +100 kJ mol<sup>-1</sup>  
(B) -400 kJ mol<sup>-1</sup> (D) +200 kJ mol<sup>-1</sup>

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol<sup>-1</sup>; O 16.0 g mol<sup>-1</sup>]

- (A) 88 kJ (B) 109 kJ  
(C) 175 kJ (D) 284 kJ

88. Calculate the value of  $\Delta H$  (in kJ mol<sup>-1</sup>) for the reaction



Bond Energies (kJ mol <sup>-1</sup> )	
H-H	435
N=N	945 (in N <sub>2</sub> )
N-H	389

- (A) 2340 kJ of heat absorbed  
(B) 213 kJ of heat absorbed  
(C) 2340 kJ of heat evolved  
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

(A) dissolving sugar in water

(B) expanding a gas

(C) evaporating a liquid

(D) freezing water

92. When Al<sub>2</sub>O<sub>3</sub>(s) is formed from the elements at standard conditions, the values of  $\Delta H^\circ$  and  $\Delta G^\circ$  at 298 K are -1676 kJ mol<sup>-1</sup> and -1577 kJ mol<sup>-1</sup>, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3  
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

(A)  $\Delta H$ ,  $\Delta S$ , and  $\Delta G$  are positive at all temperatures.

(B)  $\Delta H$  and  $\Delta S$  are positive.

(C)  $\Delta G$  is negative at low temperatures, positive at high temperatures.

(D)  $\Delta H = \Delta S$

94. A particular chemical reaction has a negative  $\Delta H$  and negative  $\Delta S$ . Which statement is correct?

(A) The reaction is spontaneous at all temperatures.

(B) The reaction is nonspontaneous at all temperatures.

(C) The reaction becomes spontaneous as temperature increases.

(D) The reaction becomes spontaneous as temperature decreases.

# American Chemical Society Final Exam Study Guide

**Camilla Rothe**



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