

- \_\_\_\_\_ 1. A chemical reaction is a process in which  
 a. products change into reactants      c. substances change state  
 b. the law of conservation of mass applies      d. all of these
- \_\_\_\_\_ 2. During a chemical reaction,  
 a. new elements are produced      c. atoms are destroyed  
 b. atoms are rearranged      d. elements are destroyed
- \_\_\_\_\_ 3. An equation is balanced by  
 a. changing subscripts      c. erasing elements as necessary  
 b. adding coefficients      d. adding elements as necessary
- \_\_\_\_\_ 4. An atom's ability to undergo chemical reactions is determined by its  
 a. protons    b. innermost electrons    c. neutrons    d. outermost electrons
- \_\_\_\_\_ 5. What are the reactants in the following chemical equation:  

$$\text{Zn} + \text{CuSO}_4 \longrightarrow \text{ZnSO}_4 + \text{Cu}$$
 a. zinc and copper      c. zinc and copper (II) sulfate  
 b. zinc sulfate and copper      d. only zinc
- \_\_\_\_\_ 6. What are the products in the above equation?  
 a. zinc and copper      c. zinc and copper (II) sulfate  
 b. zinc sulfate and copper      d. only zinc

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.



# Balancing Chemical Equations Multiple Choice Test

**James T. Shipman, Jerry D.  
Wilson, Aaron W. Todd**



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