

Vol. of 1 mol of ideal gas at STP/NTP = 22.4 L

उत्स नान एरिया
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoga ans me.

* Significant fig
- 0 to the left of the 1st non 0 digit are sig. not sig.
eg 0.0012 => 2
- Terminal 0 are not sig. if there is no decimal pt.
eg 1000 => 1 fig.
- 0 at end is sig. if there is decimal pt.
- Exponential term are not considered sig.

अंश - 3/131
axb=c { add/subtract is least amount }
3 same sabse kam fig hoga vo sig. fig ans me hoga

* LAW OF MASS CONSV
- In a reactⁿ -> actual mass of reactant = mass of prod.
Rakme zitna use hua.
- Exceptⁿ -> Nuclear Rxn (E=mc²)
- PppWedⁿ -> Balanced Rxn

MOLE CONCEPT

* LAW OF CONST PROPORTⁿ
- Elements in a comp are in fixed ratio by mass.
eg CO₂ 12 32g => 3:8
% of C = $\frac{3}{11} \times 100$

Q. cal. the mass of C req to form 132g of CO₂.
=> C + 2O -> CO₂
12g 32g 44g
44g CO₂ ke lie 12g C
∴ 132g ke lie = $\frac{12}{44} \times 132 = 36g$

* Law of Multiple Propⁿ
- elements should be same or same fixed mass
eg CO, CO₂ 1:2 ratio
Other eg. N₂O, N₂O₂, N₂O₅
exceptⁿ -> N₂O₃

* Average no. of atoms
- Same (P, T, R, U) of molecules of gases some (max eg! not of matter)
X₂, X₃, X₄
R : R : R molecules
2 : 3 : 4 atoms [different]

* Gay Lussac's law
- CONDINS: +T, P same
- at least 2 gases.
H₂ + Cl₂ -> 2HCl
1 : 1 : 2
- Vol. R₂ U₂

* DALTON'S ATOMIC THEORY
- Atoms -> Neither created nor destroyed

* MOLE CONCEPT
- Atomic No. (P) → mass no. (P+N)
eg $\frac{A}{Z} X$
- Mass of Proton = 1.67×10^{-27} kg
- Neutron = 1.67×10^{-27} kg
- Electron = 9.1×10^{-31} kg
- MP ≈ 1837 m_e

* Atomic mass unit (amu)
1 amu = $\frac{1}{12}$ Mass of ¹²C
1 amu = 1.67×10^{-27} kg
eg Mass of ¹²C = 12 amu
- Relative atomic mass
- Unitless
- eg. mass. A.M. of C = 12

* Molecular Mass
- H₂SO₄ = 98 amu
- Mass of 1 molecule of H₂SO₄
- FORMULA MASS: for ionic comp.

* MOLE
- Amount of substance.
- 1 mole = No particles (6.022 × 10²³)
1 amu < 1.67 × 10⁻²⁷ kg
Na = 6.022 × 10²³
1 amu y Na = 1g

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