

Molality is number of moles of solute present in 1 kg of solvent.

e.g. if 80 g NaOH is present in 1000 g water then the molality will be calculated number of

$$\text{moles of NaOH} = \frac{\text{Amount}}{\text{Molar Mass}}$$

$$= \frac{80}{40} = 2$$

Here 2 moles of NaOH are in 1000 g (or 1 kg)

Solvent so the molality is 2 molal

$$\text{Formulae of molality} = \frac{W_{\text{Solute}}}{M_{\text{Solute}} \times W_{\text{solvent}} (\text{kg})}$$

$$= \frac{W_{\text{Solute}} \times 1000}{M_{\text{Solute}} \times W_{\text{solvent}} (g)}$$

Mol fraction: Fraction of moles of a particular component in mixture.

e.g. In mixture of He(2 mol) and Ne(3 mol) the mol fraction of He =  $\frac{2}{2+3} = \frac{2}{5} = 0.4$

the mol fraction of Ne =  $\frac{3}{2+3} = \frac{3}{5} = 0.6$

Formula for mol fraction of component A

$$\left[ X_A = \frac{n_A}{\sum_{i=1}^n n_i} = \frac{\text{number of moles of A}}{\text{Total number of moles of all species}} \right]$$

# Mole Fraction Problems And Solutions

**Oswaal Editorial Board**



## **Mole Fraction Problems And Solutions:**

## Decoding **Mole Fraction Problems And Solutions**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Mole Fraction Problems And Solutions**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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