

Chapter 5 Chapter Review Question Answers

- Describe the scientists' contributions:
 - Cannizzaro—developed a standard method for measuring atomic masses, which allowed chemists to search for periodic trends among elements.
 - Mendeleev—organized elements according to increasing atomic mass and noticed that similar properties appeared periodically.
 - Moseley—discovered that atomic number, not atomic mass, should be the basis of organizing elements on the periodic table.
- The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- Groups of elements exhibit similar chemical and physical properties and behaviors.
- Electron configurations:
 - Generally, the configurations of the outermost electrons shells of elements within the same group are the same. There are a number of exceptions to this rule, however, among the transition elements.
 - Their outer shells are completely filled.
- The length of a period is determined by the total number of electrons that can fill the outer sublevels of the elements of that period.
- An element's period corresponds to its highest occupied main energy level. (s & p)
- Blocks & Groups
 - The type of sublevel being filled in successive elements of that block.
 - s-block: Groups 1 & 2; p-block: Groups 13–18; d-block: Groups 3–12; f-block: No group #s, as they are between groups 3 & 4, but period 6 elements are in the Lanthanide series, and period 7 elements are in the Actinide series.
- Alkali Metals
 - The Group 1 elements, except Hydrogen.
 - Alkali metals are extremely reactive, they react violently with water, they have low densities, they are soft enough to be cut with a knife, and freshly cut alkali metals are very shiny silver colored metals.
- Alkaline Metals
 - The Group 2 elements.
 - They are slightly less reactive than Group 1 elements; they are harder, duller, denser than Group 1 elements. They also have slightly higher melting points than the Group 1 elements.
- Group Configurations:
 - Transition Metals (d-block)
 - Group 3: $ns^2 (n-1)d^1$
 - Group 4: $ns^2 (n-1)d^2$
 - Group 5: $ns^2 (n-1)d^3$
 - Group 6: $ns^2 (n-1)d^4$ {actually is $ns^1 (n-1)d^5$ for Cr & Mo}
 - Group 7: $ns^2 (n-1)d^5$
 - Group 8: $ns^2 (n-1)d^6$
 - Group 9: $ns^2 (n-1)d^7$
 - Group 10: $ns^2 (n-1)d^8$
 - Group 11: $ns^2 (n-1)d^9$ {actually is $ns^1 (n-1)d^{10}$ for all group 11 elements}
 - Group 12: $ns^2 (n-1)d^{10}$
 - The group number is the sum of the s & d electrons.
- The transition elements.
- p-block
 - The p-block consists of nonmetals at the right, metalloids in the middle along the staircase, and metals at the left.
 - The p-block metals are generally harder and denser than the s-block metals, but softer and less dense than the d-block metals.
- Halogens
 - The Group 17 elements.

Modern Chemistry Chapter 5 Review Answer Key

Brendan G. Carr



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Foundations of College Chemistry, International Adaptation Morris Hein, Susan Arena, Cary Willard, 2023-05-16
Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions including clear explanations and step by step problem solving. The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans rules, collision number and mean free path, abnormal molecular masses and van t Hoff factor and applications of radioactivity.

Medicinal Chemistry Norma K Dunlap, Donna M Huryn, 2025-11-28
This book provides a concise introduction to the medicinal chemistry of therapeutically active compounds. Written in an accessible style, it offers a brief overview of the drug development process from discovery to approved drug. Emphasizing medicinal chemistry strategies, the content is organized by target as well as therapeutic area, drug classes. It covers existing drugs and the discovery process that led to their development. Aimed at students and professionals interested in drug discovery, Medicinal Chemistry requires little prior knowledge, making it ideal for classroom and training seminar use. Written by experts in the field, this outstanding new edition explains medicinal chemistry on chemical mechanism terms, allowing for a more interesting and user friendly text. Includes valuable insights toward the various pathways taken by the pharmaceutical industry in drug discoveries. Enhanced text includes questions raised and suggestions made from students in the authors medicinal chemistry class. Content provided to enhance the text includes a solutions manual, PowerPoint slides for lecture outlines, and especially for the earlier chapters, a selection of videos. An accompanying website will be updated to include additional material: 1. Journal Club exercises, current journal articles with guiding questions for student analysis; two per chapter; 2. additional case studies; 3. organic synthesis analysis, selected syntheses, both discovery and process will be discussed; and 4. updates on current advances in the field of medicinal chemistry. The book benefits a diverse audience including upper level undergraduates and graduates who are studying in the fields of medicinal chemistry and drug discovery, as well as scientists working in the pharmaceutical industry.

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998
Organic Synthesis Michael Smith, 2011-07-12
A reactions oriented course is a staple of most graduate organic programs and synthesis is taught either as a part of that course or as a special topic. Ideally, the incoming student is an organic major who has a good working knowledge of basic reactions, stereochemistry, and conformational principles. In fact, however, many often most of the students in a first year graduate level

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Introduction to Modern Chemistry: Student Manual Eugene Meyer,1979 **The Chautauquan** ,1885 *The Athenaeum* ,1871 **Athenaeum and Literary Chronicle** ,1863 **General Chemistry** Brady,1990-01-02 *Saturday Review* ,1883 The Saturday Review of Politics, Literature, Science and Art ,1860 **Machinery** Fred Herbert Colvin,Lester Gray French,1914 Foundations of College Chemistry, Alternate Morris Hein,Susan Arena,2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis *Water and Gas Review* ,1914 *Books in Print Supplement* ,2002

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