

experiment date(s):

Experiment 6 - Analysis of an Unknown Chloride

Goal: Determine the % chloride (Cl) in an unknown sample.

In the experiment, we will titrate with silver nitrate (AgNO_3) that has a known concentration.

net ionic equation: $\text{Ag}^+_{(aq)} + \text{Cl}^-_{(aq)} \longrightarrow \text{AgCl}_{(s)}$ white precipitate

When Cl^- runs out, the reaction is complete; this is called the **end point**. We will know the end point is reached because CrO_4^{2-} (chromate) is added to the chloride solution. It will also form a precipitate with Ag^+ , but only after all the Cl^- is used up

net ionic equation: $2 \text{Ag}^+_{(aq)} + \text{CrO}_4^{2-}_{(aq)} \longrightarrow \text{Ag}_2\text{CrO}_4_{(s)}$ red precipitate

The chromate ion, yellow in color, will make the titration appear to form yellow precipitate. We will stop adding the silver solution when the chloride solution begins to turn red.

ex. A solid chloride sample weighing 0.3361 g requires 61.46 mL of 0.06258 M AgNO_3 to reach the end point.

(a) How many mol of Cl^- are present in the sample?

$$\frac{0.06146 \text{ L AgNO}_3}{1 \text{ L AgNO}_3} \left(\frac{0.06258 \text{ mol AgNO}_3}{1 \text{ L AgNO}_3} \right) \left(\frac{1 \text{ mol Cl}^-}{1 \text{ mol AgNO}_3} \right) = 3.846 \times 10^{-3} \text{ mol Cl}^-$$

(b) How many g of Cl^- are present in the sample?

$$\frac{3.846 \times 10^{-3} \text{ mol Cl}^-}{1 \text{ mol Cl}^-} \left(\frac{35.45 \text{ g Cl}}{1 \text{ mol Cl}} \right) = 0.1363 \text{ g Cl}^-$$

(c) What is the mass % of Cl^- in the sample?

$$\frac{\text{g Cl}^-}{\text{g sample}} \times 100 = \frac{0.1363 \text{ g}}{0.3361 \text{ g}} \times 100 = 40.55 \%$$

Procedure: omit back titration with NaCl

Waste disposal: Dispose of all liquids and rinse all solids into the Heavy Metal Inorganic waste container.

ASA: all

Analysis Of An Unknown Chloride Answers

J Dewey



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