

Lecture 18 The End-behavior of Polynomial Functions

Notation: $x \rightarrow \infty$ means x becomes large, positive

$x \rightarrow -\infty$ means x becomes large, negative (i.e. $|x|$ is large and $x < 0$)

When $x \rightarrow \infty$, $f(x) = a_n x^n + a_{n-1} x^{n-1} + \dots + a_1 x + a_0$ behaves as the leading term $a_n x^n$.

(Other terms are much smaller because their exponents are smaller.)

$$\begin{aligned} \text{Ex: } f(x) = 2x^3 + x^2 + 3x + 1 &= 2x^3 \left(1 + \frac{x^2}{2x^3} + \frac{3x}{2x^3} + \frac{1}{2x^3} \right) \\ &= 2x^3 \left(1 + \frac{1}{2x} + \frac{3}{2x^2} + \frac{1}{2x^3} \right) \end{aligned}$$

So, if x is large, then $\frac{1}{2x}$, $\frac{3}{2x^2}$, $\frac{1}{2x^3}$ are small and $f(x) \approx 2x^3$ is the leading term.



So, we have four types of the end-behavior for polynomial functions:

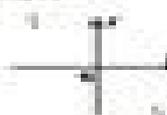


$$\text{Ex: } g(x) = -2x^3 + 3x^2 - 10x + 1$$

If $x \rightarrow \infty$, then $g(x) \rightarrow -\infty$

If $x \rightarrow -\infty$, then $g(x) \rightarrow \infty$

(End-behavior:



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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling

point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube end c

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b

had a melting point range of 122.6–124.5 °C thus

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Melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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Exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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Laboratory partners Samantha and Mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that need to be determined by physical properties such as boiling point melting point density solubility etc in this experiment i am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a bunsen burner boiling

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there are a variety of methods by which a sample s boiling point can be determined including distillation reflux and by using a thiele tube the most straightforward method uses a thiele tube and has the advantage of using less than 0 5ml 0 5 ml of material

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