

The background of the cover is a dark, almost black, field filled with a complex, fractal-like pattern of thin, glowing blue and white lines. These lines form a dense, swirling structure that resembles a mathematical curve or a complex network, with many lines radiating from a central point and curving away. The overall effect is one of intricate, organic complexity.

Discrete Mathematics

Seventh Edition

Richard Johnsonbaugh

Discrete Mathematics Johnsonbaugh 7th Edition Pdf

Michael Brown



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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} \times 18 \text{ g } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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you have 0.216232 grams/mole \times 21.3 grams = 0.66 moles \times 32 g/mol

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained

calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0

g Al_2O_3 \times 1 mol Al_2O_3 \times 2 mol Al_2O_3 \times 101.964 g Al_2O_3 = 302.3 g Al_2O_3 26.982 g Al_2O_3 \times 4 mol Al_2O_3 = 107.728 g Al_2O_3 c determine the percent yield

yield actual $\times 100$ = $\frac{260.0}{302.3} \times 100 = 86.01\%$ theoretical 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4.33 g of H₂ are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is

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reaction

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