

ZOOL Chapter 2: The Chemical Level of Organization

2.1 How Matter is Organized

- Chemistry – science of structure and interactions of matter
 - o Matter – anything that occupies space and mass
 - o Mass – the amount of matter in an object
 - o Weight – force of gravity acting on matter

Chemical Elements

- Three types of matter: solids, liquids, and gasses
- Chemical elements
 - o Make up all forms of living matter
 - o Chemical symbol: N, O, Au, Fe
- Major elements: O, C, H, N
 - o Makes up 96% of body's mass
- Lesser elements (8)
- Trace elements (14)
 - o Exist in trace amounts

Structure of Atoms

- Atoms – smallest unit of matter
- Subatomic particles:
 - o Electrons
 - o Neutrons
 - o Protons

Atomic Mass and Number

- Atomic number = # of protons
- Mass number = # of protons + neutrons
- Isotopes are atoms of an element that have different amounts of neutrons
 - o Radioactive isotopes are unstable
 - As they decay, they emit radiation

Atomic Mass

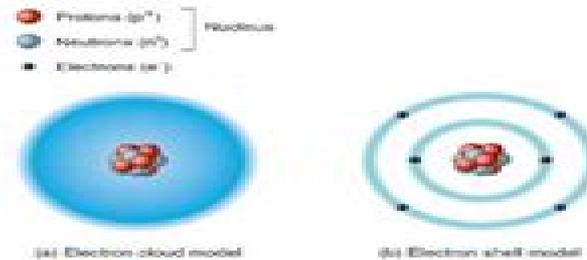
- Atomic mass unit
- A neutron has 1.008 amu
- Protons have 1.007 amu

Ions, Molecules, and Compounds

- Ion is an atom with a positive or negative charge
- Ionization – process of giving up or gaining electrons
- Molecules – two atoms SHARE electrons
- Compounds – substance that contains two or more atoms of different elements
- Free radical – atom or group of atoms with an unpaired electron

2.2 Chemical Bonds

- Chemical bonds – the forces that holds atoms of a molecule/compound together
- Valence shell – outermost energy level
- Octet rule – electrons seek to fill out all energy level



Anatomy And Physiology Chapter 2

Mary Miller, Crispian Scully



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