



7. What is the maximum mass of aluminum hydroxide that can be produced when 25.0 mL of 0.248 M sodium hydroxide solution are mixed with 25.0 mL of 0.189 M aluminum nitrate solution?

Find mol of each reactant

$$0.0250 \text{ L} \times 0.248 \frac{\text{mol}}{\text{L}} = 6.2 \times 10^{-3} \text{ mol NaOH}$$

$$0.0250 \text{ L} \times 0.189 \frac{\text{mol}}{\text{L}} = 4.73 \times 10^{-3} \text{ mol Al}(\text{NO}_3)_3$$

$$6.20 \times 10^{-3} \text{ mol NaOH} \times \frac{1 \text{ mol Al}(\text{OH})_3}{3 \text{ mol NaOH}} \times \frac{78.0 \text{ g Al}(\text{OH})_3}{\text{mol}} =$$

$$\boxed{0.161 \text{ g Al}(\text{OH})_3}$$

need 3 times as much NaOH as  $\text{Al}(\text{NO}_3)_3$  but we only have 1.3 x as much NaOH is limiting

8. Indicate which of the following are redox reactions. For those that are redox reactions, label the reactant that is the oxidizing agent and the reactant that is the reducing agent. You do not have to balance these reactions!



Redox agent agent  
red. agent ox. agent

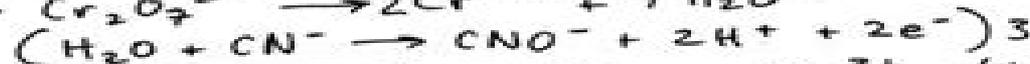


Not redox

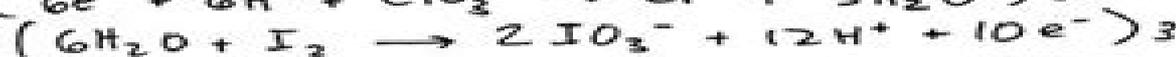


red agent ox agent

9. Aqueous dichromate ion reacts with aqueous cyanide ion ( $\text{CN}^-$ ) to produce chromium (III) ion and aqueous cyanate ( $\text{CNO}^-$ ) ion. Balance the equation using the half-reaction method, assuming the reaction occurs in acidic solution.



10. Balance the following redox reaction in basic solution.



# Answers To Modern Chemistry Homework Chapter5

**Christian G. Meyer**



## Answers To Modern Chemistry Homework Chapter5:

Fundamentals of Chemistry: A Modern Introduction (1966) Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro s Law and Graham s Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry

**Children's Books in Print** R R Bowker Publishing, Bowker, 1999-12

Foundations of Inorganic Chemistry Gary Wulfsberg, 2017-11-02 Foundations of Inorganic Chemistry by Gary Wulfsberg is our newest entry into the field of Inorganic Chemistry textbooks designed uniquely for a one semester stand alone course or to be used in a full year inorganic sequence Foundations of Inorganic Chemistry by Gary Wulfsberg is our newest entry into the field of Inorganic Chemistry textbooks designed uniquely for a one semester stand alone course or to be used in a full year inorganic sequence By covering virtually every topic in the test from the 2016 ACS Exams Institute this book will prepare your students for success The new book combines careful pedagogy clear writing beautifully rendered two color art and solved examples with a broad array of original chapter ending exercises It assumes a background in General Chemistry but reviews key concepts and also assumes enrollment in a Foundations of Organic Chemistry course Symmetry and molecular orbital theory are introduced after the student has developed an understanding of fundamental trends in chemical properties and reactions across the periodic table which allows MO theory to be more broadly applied in subsequent chapters Use of this text is expected to increase student enrollment and build students appreciation of the central role of inorganic chemistry in any allied field Key Features Over 900 end of chapter exercises half answered in the back of the book Over 180 worked examples Optional experiments demos Clearly cited connections to other areas in chemistry and chemical sciences Chapter opening biographical vignettes of noted scientists in Inorganic Chemistry Optional General Chemistry review sections Originally rendered two color illustrations throughout

**Books in Print Supplement**, 2002

**Basic Chemistry Concepts and Exercises** John Kenkel, 2011-07-08 Chemistry can be a daunting subject for the uninitiated and all too often introductory

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**Resources in Education**, 1968 *Research in Education*, 1968-07 *Otto E. Miller, Plaintiff-Respondent, Against Fred W. Smythe, Defendant-Appellant*, *Chemistry in the Modern World* Frank L. Wiseman, 1985 Chapter Resource 13 Theory/Evolution Biology Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004 *Quantum Mechanics* Claude Cohen-Tannoudji, Bernard Diu, Franck Laloe, 2020 **World History**, 2000 **Ate Wh** Holt Rinehart & Winston, 2000

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