

KEY

## Chemistry 12

### Worksheet 1-3 - Reaction Mechanisms

1. It is known that compounds called *chlorofluorocarbons* (C.F.C.s) (eg.  $\text{CFCl}_3$ ) will break up in the presence of ultraviolet radiation, such as found in the upper atmosphere, forming single chlorine atoms:



The Cl atoms then react with Ozone ( $\text{O}_3$ ) as outlined in the following mechanism.



- a) Write the equation for the *overall reaction*. (Using steps 1 and 2)



- b) What is the *catalyst* in this reaction?

Cl

- c) Identify an *intermediate* in this reaction

ClO

- d) Explain how a *small* amount of chlorofluorocarbons can destroy a *large* amount of ozone.

When a Cl atom destroys an  $\text{O}_3$  molecule (step 1), it is regenerated (step 2) so it can destroy more

- e) What breaks the bond in the  $\text{CFCl}_3$  and releases the free Cl atom?

a photon ( $h\nu$ ) of ultraviolet radiation

2. Given the following mechanism, answer the questions below:



- a) Give the equation for the *overall reaction*.



- b) What could the *catalyst* be in this mechanism?

NO

- c) What is an *intermediate* in this mechanism?

$\text{NO}_2$

# Answers To Unit 5 Post Test Chemistry

**Marion Castellucci**



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