

2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)



b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of H<sub>2</sub>O, which is the limiting reactant?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.0 \text{ g H}_2\text{O}$$

There is less H<sub>2</sub>O (12.0g) in our initial amount so  
H<sub>2</sub>O is the L.R

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} = 1.3 \text{ mol H}_2\text{O}$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.666 \text{ mol H}_2\text{O}$$

$$1.3 \text{ mol H}_2\text{O} - 0.666 \text{ mol H}_2\text{O} = 0.634 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Mg}}{2 \text{ mol H}_2\text{O}} = \boxed{0.317 \text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Mg}(\text{OH})_2}{2 \text{ mol H}_2\text{O}} \times \frac{58.32 \text{ g Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{OH})_2} = 19.4 \text{ g Mg}(\text{OH})_2$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} = 6.28 \text{ g H}_2$$

e. You perform this reaction, and obtain 16.0 g Mg(OH)<sub>2</sub>. What is your percent yield?

$$\% \text{ yield} = \frac{\text{actual} \times 100}{\text{theoretical}}$$

$$\frac{16.0 \text{ g Mg}(\text{OH})_2}{19.4 \text{ g Mg}(\text{OH})_2} \times 100 = 82.5\%$$

# Cookie Stoichiometry Answers

**Victor M. Corman**



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## **Table of Contents Cookie Stoichiometry Answers**

1. Understanding the eBook Cookie Stoichiometry Answers
  - The Rise of Digital Reading Cookie Stoichiometry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Cookie Stoichiometry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Cookie Stoichiometry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Cookie Stoichiometry Answers
  - Personalized Recommendations
  - Cookie Stoichiometry Answers User Reviews and Ratings
  - Cookie Stoichiometry Answers and Bestseller Lists
5. Accessing Cookie Stoichiometry Answers Free and Paid eBooks

- Cookie Stoichiometry Answers Public Domain eBooks
  - Cookie Stoichiometry Answers eBook Subscription Services
  - Cookie Stoichiometry Answers Budget-Friendly Options
6. Navigating Cookie Stoichiometry Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Cookie Stoichiometry Answers Compatibility with Devices
    - Cookie Stoichiometry Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Cookie Stoichiometry Answers
    - Highlighting and Note-Taking Cookie Stoichiometry Answers
    - Interactive Elements Cookie Stoichiometry Answers
  8. Staying Engaged with Cookie Stoichiometry Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Cookie Stoichiometry Answers
  9. Balancing eBooks and Physical Books Cookie Stoichiometry Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Cookie Stoichiometry Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Cookie Stoichiometry Answers
    - Setting Reading Goals Cookie Stoichiometry Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Cookie Stoichiometry Answers
    - Fact-Checking eBook Content of Cookie Stoichiometry Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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