

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, KClO_3 . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{22} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(III) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 15.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. (Write a balanced equation first.)

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of nitrogen dioxide. Assume all gases are at the same conditions of temperature and pressure.



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Sabine Zange

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1. Understanding the eBook Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - The Rise of Digital Reading Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Advantages of eBooks Over Traditional Books
2. Identifying Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - User-Friendly Interface
4. Exploring eBook Recommendations from Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Personalized Recommendations
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf User Reviews and Ratings
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf and Bestseller Lists

5. Accessing Core Teaching Resources Chemistry Answers Chapter 9 Pdf Free and Paid eBooks
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf Public Domain eBooks
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf eBook Subscription Services
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf Budget-Friendly Options
6. Navigating Core Teaching Resources Chemistry Answers Chapter 9 Pdf eBook Formats
 - ePub, PDF, MOBI, and More
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf Compatibility with Devices
 - Core Teaching Resources Chemistry Answers Chapter 9 Pdf Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Highlighting and Note-Taking Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Interactive Elements Core Teaching Resources Chemistry Answers Chapter 9 Pdf
8. Staying Engaged with Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Core Teaching Resources Chemistry Answers Chapter 9 Pdf
9. Balancing eBooks and Physical Books Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Core Teaching Resources Chemistry Answers Chapter 9 Pdf
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Setting Reading Goals Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Fact-Checking eBook Content of Core Teaching Resources Chemistry Answers Chapter 9 Pdf
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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