

# THERMODYNAMICS

## PROPERTIES OF SINGLE-COMPONENT SYSTEMS

### Nomenclature

1. Intensive properties are independent of mass.
2. Extensive properties are proportional to mass.
3. Specific properties are lowercase (extensive/mass).

### State Functions (properties)

Absolute Pressure,  $P$  (lb/in<sup>2</sup> or Pa)

Absolute Temperature,  $T$  (°R or K)

Volume,  $V$  (ft<sup>3</sup> or m<sup>3</sup>)

Specific Volume,  $v = V/m$  (ft<sup>3</sup>/lbm or m<sup>3</sup>/kg)

Internal Energy,  $U$  (Btu or kJ)

Specific Internal Energy,  $u = U/m$  (usually in Btu/lbm or kJ/kg)

Enthalpy,  $H$  (Btu or kJ)

Specific Enthalpy,  $h = u + Pv = H/m$  (usually in Btu/lbm or kJ/kg)

Entropy,  $S$  (Btu/°R or kJ/K)

Specific Entropy,  $s = S/m$  [Btu/(lbm·°R) or kJ/(kg·K)]

Gibbs Free Energy,  $g = h - Ts$  (usually in Btu/lbm or kJ/kg)

Helmholtz Free Energy,  $a = u - Ts$  (usually in Btu/lbm or kJ/kg)

Heat Capacity at Constant Pressure,  $c_p = \left(\frac{\partial h}{\partial T}\right)_p$

Heat Capacity at Constant Volume,  $c_v = \left(\frac{\partial u}{\partial T}\right)_v$

Quality  $x$  (applies to liquid-vapor systems at saturation) is defined as the mass fraction of the vapor phase:

$$x = m_g / (m_g + m_l), \text{ where}$$

$m_g$  = mass of vapor, and

$m_l$  = mass of liquid.

Specific volume of a two-phase system can be written:

$$v = xv_g + (1-x)v_f \text{ or } v = v_f + xv_{fg} \text{ where}$$

$v_f$  = specific volume of saturated liquid,

$v_g$  = specific volume of saturated vapor, and

$$v_{fg} = \text{specific volume change upon vaporization,} \\ = v_g - v_f$$

Similar expressions exist for  $u$ ,  $h$ , and  $s$ :

$$u = xu_g + (1-x)u_f \text{ or } u = u_f + xu_{fg}$$

$$h = xh_g + (1-x)h_f \text{ or } h = h_f + xh_{fg}$$

$$s = xs_g + (1-x)s_f \text{ or } s = s_f + xs_{fg}$$

For a simple substance, specification of any two intensive, independent properties is sufficient to fix all the rest.

For an ideal gas,  $Pv = RT$  or  $PV = mRT$ , and

$$P_1 v_1 / T_1 = P_2 v_2 / T_2, \text{ where}$$

$P$  = pressure,

$v$  = specific volume,

$m$  = mass of gas,

$R$  = gas constant, and

$T$  = absolute temperature.

$V$  = volume

$R$  is specific to each gas but can be found from

$$R = \frac{\bar{R}}{(\text{mol. wt})}, \text{ where}$$

$\bar{R}$  = the universal gas constant

$$= 1,545 \text{ ft}\cdot\text{lb}/(\text{lbmol}\cdot^\circ\text{R}) = 8,314 \text{ J}/(\text{kmol}\cdot\text{K}).$$

For ideal gases,  $c_p - c_v = R$

Also, for ideal gases:

$$\left(\frac{\partial h}{\partial P}\right)_T = 0 \quad \left(\frac{\partial u}{\partial v}\right)_T = 0$$

For cold air standard, heat capacities are assumed to be constant at their room temperature values. In that case, the following are true:

$$\Delta u = c_v \Delta T, \quad \Delta h = c_p \Delta T$$

$$\Delta s = c_p \ln(T_2/T_1) - R \ln(P_2/P_1); \text{ and}$$

$$\Delta s = c_v \ln(T_2/T_1) + R \ln(v_2/v_1).$$

For heat capacities that are temperature dependent, the value to be used in the above equations for  $\Delta h$  is known as the mean heat capacity ( $\bar{c}_p$ ) and is given by

$$\bar{c}_p = \frac{\int_{T_1}^{T_2} c_p dT}{T_2 - T_1}$$

Also, for constant entropy processes:

$$\frac{P_2}{P_1} = \left(\frac{v_1}{v_2}\right)^k; \quad \frac{T_2}{T_1} = \left(\frac{P_2}{P_1}\right)^{\frac{k-1}{k}}$$

$$\frac{T_2}{T_1} = \left(\frac{v_1}{v_2}\right)^{k-1}, \text{ where } k = c_p/c_v$$

For real gases, several equations of state are available; one such equation is the van der Waals equation with constants based on the critical point:

$$\left(P + \frac{a}{v^2}\right)(v - b) = RT$$

$$\text{where } a = \left(\frac{27}{64}\right)\left(\frac{R^2 T_c^2}{P_c}\right), \quad b = \frac{RT_c}{8P_c}$$

where  $P_c$  and  $T_c$  are the pressure and temperature at the critical point, respectively, and  $v$  is the molar specific volume.

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