

Chapter Test B

Chapter: Chemical Equations and Reactions

Match a name space (possibly writing the name of the react or product molecule) completed with reactants/products atoms/molecules.

- The product(s) of a chemical reaction will be represented by a double arrow pointing to the right and the reactant(s) will be represented by a double arrow pointing to the left.
 - reactants**
 - products
 - solvents
- An ionic chemical equation is represented by adding the
 - charges
 - structure of the particles
 - number of particles
- An ionic equation shows a reaction by adding particles that are in the
 - state
 - aqueous or solid
 - solvent
- The activity for precipitation reactions on pages 100-101 is that
 - soluble ionic compounds react and form an insoluble precipitate
 - insoluble ionic compounds react and form a soluble precipitate
 - soluble ionic compounds react
 - soluble ionic compounds react and form an insoluble precipitate**
- The coefficients are balanced equations
 - indicate the number of moles of each substance that react**
 - show the number of grams of each substance that react
 - show the number of moles of the substances
 - show the relative proportions of each atom
- If you're to balance the reaction $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ by using the variable, then
 - X will replace each of the coefficients**
 - X will replace each of the subscripts
 - X will replace each of the formulas
 - X will replace each of the atoms

Chapter 8 Test B

Christian G. Meyer



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