

Name: \_\_\_\_\_

Number: \_\_\_\_\_

19. Atmospheric pressure at elevations of 1000 ft average about 0.70 atm. Consider this pressure as equilibrium of nitrogen, oxygen and argon and assume each behaves as an independent gas. What is the partial pressure of  $O_2$  at this elevation?

Answer: Partial pressures of gases in a mixture are proportional to their mole fractions. The partial pressure of  $O_2$  is proportional to its mole fraction in the mixture. The mole fraction of  $O_2$  in air is 0.21. The partial pressure of  $O_2$  is  $0.21 \times 0.70 \text{ atm} = 0.147 \text{ atm}$ .

20. Suppose a 1.00 liter balloon inflated with a volume of 2.0 L, from a cylinder, during the process at 1.0 atm, has a depth of 200 m, where the pressure is 3.0 atm. What would happen to the volume of the balloon? What if the end of the balloon was in a tank of water that went to the surface and was released to another balloon? What would happen to the volume of the balloon?

21. A volume of gas is confined in a cylinder with a freely moving piston, and held at 25°C. What happens to the pressure in the cylinder if the volume is doubled? What happens to the volume of the cylinder if the pressure is doubled? What happens to the volume of the cylinder if the pressure is halved?

- (a) The pressure will increase to half that the volume will double.
- (b) The pressure will increase to half the volume will double.
- (c) The pressure will increase to half that the volume will increase four times.
- (d) The pressure will increase four times the volume will double.

22. Use Charles's law to compare the volume of a gas at 25°C and 1.00 atm to the volume of the gas at 100°C.

$T_1$	$T_2$	$V_1$	$V_2$
25°C	100°C	1.00 L	1.36 L
100°C	25°C	1.36 L	1.00 L

# Chem 1050 Homework Exam 1 Assignment Solutions

**Jianjun Gao**



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