

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- | | |
|-----------------------------|---|
| a. Sulfur hexafluoride | SF_6 |
| b. Lithium nitride | Li_3N |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$ |
| d. Tin (II) chloride | SnCl_2 |
| e. Ammonium acetate | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$) |
| f. Mercury (I) chloride | Hg_2Cl_2 |
| g. Potassium bromate | KBrO_3 |
| h. Hydrosulfuric acid | $\text{H}_2\text{S(aq)}$ |
| i. Chloric acid | HClO_3 |
| j. Sulfurous acid | H_2SO_3 |

Name each of the following compounds:

- | | |
|------------------------------|-------------------------|
| a. CuSO_4 | Copper (II) sulfate |
| b. AlF_3 | aluminum fluoride |
| c. HI | hydrogen iodide |
| d. NO | nitrogen monoxide |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid |
| f. HNO_3 | nitric acid |
| g. HNO_2 | nitrous acid |
| h. NaHSO_4 | sodium hydrogen sulfate |

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Rachel Sandford



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Table of Contents Chemistry Chapter9 Review Answers

1. Understanding the eBook Chemistry Chapter9 Review Answers
 - The Rise of Digital Reading Chemistry Chapter9 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter9 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter9 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter9 Review Answers
 - Personalized Recommendations
 - Chemistry Chapter9 Review Answers User Reviews and Ratings
 - Chemistry Chapter9 Review Answers and Bestseller Lists

5. Accessing Chemistry Chapter9 Review Answers Free and Paid eBooks
 - Chemistry Chapter9 Review Answers Public Domain eBooks
 - Chemistry Chapter9 Review Answers eBook Subscription Services
 - Chemistry Chapter9 Review Answers Budget-Friendly Options
6. Navigating Chemistry Chapter9 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter9 Review Answers Compatibility with Devices
 - Chemistry Chapter9 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter9 Review Answers
 - Highlighting and Note-Taking Chemistry Chapter9 Review Answers
 - Interactive Elements Chemistry Chapter9 Review Answers
8. Staying Engaged with Chemistry Chapter9 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter9 Review Answers
9. Balancing eBooks and Physical Books Chemistry Chapter9 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter9 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter9 Review Answers
 - Setting Reading Goals Chemistry Chapter9 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter9 Review Answers
 - Fact-Checking eBook Content of Chemistry Chapter9 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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