

Name Key

1. Hydrogen, H<sub>2</sub>, is used as rocket fuel. The hydrogen is burned in oxygen, O<sub>2</sub>, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H<sub>2</sub> is burned?



$$\frac{200.0 \text{ g H}_2}{2.016 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1} = 99.2 \text{ mol H}_2$$

$$99.2 \text{ mol H}_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol H}_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?

reverse so change sign of  $\Delta H$  - that's in 1/2 quantity

$$-\frac{1}{2} \text{CO}_2(\text{g}) + \text{SO}_2(\text{g}) \rightarrow \frac{1}{2} \text{CS}_2(\text{l}) + \frac{3}{2} \text{O}_2(\text{g})$$

$$+ \frac{1077 \text{ kJ}}{2} = 538.5 \text{ kJ}$$

3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol C}_3\text{H}_6\text{O}} \times 4.5 \text{ mol C}_3\text{H}_6\text{O} = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is 0.110 cal / g · °C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot \text{C}} \right) (225 - 25 \text{ C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N<sub>2</sub>H<sub>4</sub>, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



Reverse  
doubly quantity

$$\begin{array}{r} \text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow \text{N}_2\text{H}_4(\text{l}) + \text{O}_2(\text{g}) \quad \Delta H = 622.2 \text{ kJ} \\ 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) \quad \Delta H = (2)(-285.8) \text{ kJ} \\ \hline \text{N}_2(\text{g}) + 2 \text{H}_2(\text{g}) \rightarrow \text{N}_2\text{H}_4(\text{l}) \quad \Delta H = +50.6 \text{ kJ} \end{array}$$

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## **Table of Contents Chemistry Practice Problems With Answers Activity Series**

1. Understanding the eBook Chemistry Practice Problems With Answers Activity Series
  - The Rise of Digital Reading Chemistry Practice Problems With Answers Activity Series
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Problems With Answers Activity Series
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Practice Problems With Answers Activity Series
  - User-Friendly Interface

4. Exploring eBook Recommendations from Chemistry Practice Problems With Answers Activity Series
  - Personalized Recommendations
  - Chemistry Practice Problems With Answers Activity Series User Reviews and Ratings
  - Chemistry Practice Problems With Answers Activity Series and Bestseller Lists
5. Accessing Chemistry Practice Problems With Answers Activity Series Free and Paid eBooks
  - Chemistry Practice Problems With Answers Activity Series Public Domain eBooks
  - Chemistry Practice Problems With Answers Activity Series eBook Subscription Services
  - Chemistry Practice Problems With Answers Activity Series Budget-Friendly Options
6. Navigating Chemistry Practice Problems With Answers Activity Series eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Practice Problems With Answers Activity Series Compatibility with Devices
  - Chemistry Practice Problems With Answers Activity Series Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Practice Problems With Answers Activity Series
  - Highlighting and Note-Taking Chemistry Practice Problems With Answers Activity Series
  - Interactive Elements Chemistry Practice Problems With Answers Activity Series
8. Staying Engaged with Chemistry Practice Problems With Answers Activity Series
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Practice Problems With Answers Activity Series
9. Balancing eBooks and Physical Books Chemistry Practice Problems With Answers Activity Series
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Practice Problems With Answers Activity Series
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practice Problems With Answers Activity Series
  - Setting Reading Goals Chemistry Practice Problems With Answers Activity Series
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Practice Problems With Answers Activity Series
  - Fact-Checking eBook Content of Chemistry Practice Problems With Answers Activity Series
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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