

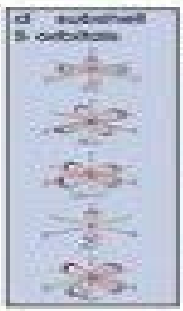
CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1** - Atomic Number
- H** - Atomic Symbol
- 1.008** - Atomic Weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For a neutral atom, $Z = +$ the number of electrons. The mass number, A , is the sum of protons and neutrons in the nucleus. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Structure, Atomic Weight, Crystal Structure of Solids, Electron Spin, and Quantum Numbers.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum numbers describe the energy, size, and shape of an atomic orbital. The principal quantum number, n , is the most important. It is the number of nodes in the radial wave function. The angular momentum quantum number, l , is the number of nodes in the angular wave function. The magnetic quantum number, m_l , is the number of nodes in the magnetic field. The spin quantum number, m_s , is the number of nodes in the spin wave function.

NUCLEAR CHEMISTRY

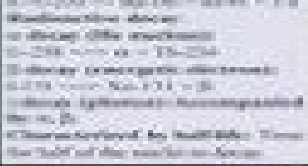
Nuclear reactions occur when nucleons (protons and neutrons) are rearranged. The energy released or absorbed is the binding energy. The mass defect is the difference between the mass of the nucleus and the sum of the masses of its constituent nucleons. The binding energy per nucleon is a measure of the stability of a nucleus. The half-life of a radioactive isotope is the time it takes for half of the nuclei to decay.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the energy levels of many-electron atoms. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy. The Hund's Rule of Maximum Multiplicity states that electrons fill degenerate orbitals singly before pairing up.

n	l	m _l	m _s	Orbital
1	0	0	±1/2	1s
2	0	0	±1/2	2s
2	1	0	±1/2	2p _z
2	1	±1	±1/2	2p _x , 2p _y
3	0	0	±1/2	3s
3	1	0	±1/2	3p _z
3	1	±1	±1/2	3p _x , 3p _y
3	2	0	±1/2	3d _{z²}
3	2	±1	±1/2	3d _{xy} , 3d _{xz} , 3d _{yz}

Atomic configuration: The arrangement of electrons in an atom. The Pauli Exclusion Principle, Hund's Rule of Maximum Multiplicity, and the Aufbau Principle are used to determine the ground state configuration of an atom. The energy levels of an atom are determined by the principal quantum number, n , and the angular momentum quantum number, l .



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Mark D. Jackson, BarCharts, Inc



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Chemistry BarCharts Inc,2012-05-31 Chemistry Mark D. Jackson,BarCharts, Inc,2001-01-24 This 4 page laminated guide is essential for students or individual looking for Chemistry knowledge The QuickStudy Chemistry guide includes information on atomic structure nuclear chemistry atomic quantum numbers orbitals electronic atoms types of matter reactions physical processes nomenclature stoichiometry chemical interactions bonding models gasesthermodynamics theories and much more QuickStudy Academic Mark Jackson (PhD and Science Professor), BarCharts best selling quick reference to chemistry has been updated and expanded in this new edition With updated content and an additional panel of information this popular guide is not only an essential companion for students in introductory chemistry courses but also a must have refresher for students in higher level courses Author Mark D Jackson PhD a scientist and university chemistry professor has a gift for making the complicated subject of chemistry interesting and easy to understand without the fluff In this new edition you will find more coverage of the subject helpful illustrations chemical problems and practical applications making this a study tool you won t want to be without **Organic Chemistry Fundamentals** Mark Jackson,2015-12-31 Quick Reference for the core essentials of a subject and class that is challenging at best and that many students struggle with In 6 laminated pages our experienced chemistry author and professor gathered key elements organized and designed to use along with your text and lectures as a review before testing or as a memory companion that keeps key answers always at your fingertips As many students have said a must have study tool Suggested uses o Quick Reference instead of digging into the textbook to find a core answer you need while studying use the guide to reinforce quickly and repeatedly o Memory refreshing your memory repeatedly is a foundation of studying have the core answers handy so you can focus on understanding the concepts o Test Prep no student should be cramming but if you are there is no better tool for that final review **Chemistry Equations and Answers** Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students Organic Chemistry Reactions BarCharts, Inc,Mark Jackson,2015-12-31 Quick Reference for the core essentials of a subject and class that is challenging at best and that many students struggle with In 6 laminated pages our experienced chemistry author and professor gathered key elements organized and designed to use along with your text and lectures as a review before testing or as a memory companion that keeps key answers always at your fingertips As many students have said a must have study tool Suggested uses o Quick Reference instead of digging into the textbook to find a core answer you need while studying use the guide to reinforce quickly and repeatedly o Memory refreshing your memory repeatedly is a foundation of studying have the core answers handy so you can focus on understanding the concepts o Test Prep no student should be cramming but if you are there is no better tool for that final review Chem Lab Basics Mark Jackson, Frank Miskevich,2018-05 With the study of molecules and how they interact with one another theory is important but the lab is where the true excitement of chemistry lies This 6 page

laminated guide is intended to serve as a reminder and quick study review for all aspects of the chemistry lab from safety to chemicals to instruments and safety again first last and always Designed to find answers fast and to be spill proof it s perfect for high school and early college students studying chemistry as well as anyone interested in learning how scientists safely study the molecules that make up everything around us 6 page laminated guide includes Lab Safety Training General Lab Guidelines Working With Chemicals Exposure to Chemicals First Aid Know Your Lab Reagents Chemical Spills Waste Management Useful Chemical Information Periodic Table Included Lab Must Knows Safe Use of Lab Equipment Data Manipulation Preparing a Solution **Inorganic Chemistry** Mark Jackson,2010-05-31 Our 3 panel 6 page guide is jam packed with chemistry information not covered in organic chemistry and biochemistry Perfect for science students who want to focus on the aspects of chemistry that go beyond what is covered in material that deals with organic topics this guide covers the key concepts principles figures and formulas that inorganic chemistry students will need to know in order to succeed Tables images and graphic elements further enhance the text *Chemistry 2* BarCharts, Inc,Jane P. Gardner,2016

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