

SEMESTER 1 FINAL REVIEW – CHEMISTRY

2015-16

Please ensure that you know how to answer these questions correctly, these are all good example problems for your final exam. Study your notes, labs, homework assignments and use the PowerPoint presentations that are online. Any of this information is fair game to be asked on the final exam. Good luck and ask questions in class during the review times if you have them. Don't wait to the last minute to start reviewing!

1. What is the number of significant figures in 1007000.0
2. What is the difference between accuracy and precision?
3. What is Avogadro's number? Why is it useful?
4. Find the molecular weight of CO_2 , NH_3 , SF_6 ?
5. Be able to use density to calculate mass and or volume for various objects. Also, understand that density can be used at times with moles and particles through the $D=m/v$ formula.
6. What is a mole? Be able to explain how the mole is a central connecting idea between mass, particles and gases @STP
7. What is the molar volume of a gas at STP?
8. What are the seven diatomic elements?
9. Define: atom, molecule, diatomic gas, monoatomic gas, element, compound, STP
10. What elements form the following bonds: ionic, covalent
11. Give an example of the following (both name and formula)
 - a. an ionic compound
 - b. a covalent compound
 - c. a metal
 - d. a nonmetal
 - e. a diatomic gas
12. Name the following:

a) NaOH	b) SO_2	c) C_2SO_4	d) HCl	e) $\text{Ca}(\text{NO}_3)_2$	f) NH_4
g) N_2O_5	h) $\text{Al}_2(\text{SO}_4)_3$	i) K_3PO_4	j) Na_2O	k) CO	l) CO_2
13. Write formulas for following
 - a) dinitrogen trioxide
 - b) potassium sulfate
 - c) hydrogen phosphate
 - d) sodium bromide
 - e) ammonium acetate
 - f) manganese IV nitrate
 - g) iron III hydroxide
 - h) aluminum carbonate
14. Write the following in correct scientific notation:

a) 476	b) 0.0000000002223	c) 3240000000000000
d) 462.9×10^{23}	e) $0.000000345 \times 10^{12}$	
f) 5677.9×10^{-12}	g) $0.000000223 \times 10^{-6}$	
15. How many oxygen atoms are there in 420 grams of carbon dioxide?
16. What is the mass of 1.2×10^{24} molecules of bromine?
17. What is the volume occupied by 243 grams of oxygen gas at STP?
18. How do you convert from:
 - a) moles \rightarrow grams
 - b) grams \rightarrow moles
 - c) moles \rightarrow molecules
 - d) moles \rightarrow liters of gas at STP
 - e) moles \rightarrow liters of solution

Chemistry Semester 1 Review Answers

Michael Seilmaier



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Table of Contents Chemistry Semester 1 Review Answers

1. Understanding the eBook Chemistry Semester 1 Review Answers
 - The Rise of Digital Reading Chemistry Semester 1 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Semester 1 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Semester 1 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Semester 1 Review Answers
 - Personalized Recommendations
 - Chemistry Semester 1 Review Answers User Reviews and Ratings
 - Chemistry Semester 1 Review Answers and Bestseller Lists
5. Accessing Chemistry Semester 1 Review Answers Free and Paid eBooks
 - Chemistry Semester 1 Review Answers Public Domain eBooks
 - Chemistry Semester 1 Review Answers eBook Subscription Services

- Chemistry Semester 1 Review Answers Budget-Friendly Options
- 6. Navigating Chemistry Semester 1 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Semester 1 Review Answers Compatibility with Devices
 - Chemistry Semester 1 Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Semester 1 Review Answers
 - Highlighting and Note-Taking Chemistry Semester 1 Review Answers
 - Interactive Elements Chemistry Semester 1 Review Answers
- 8. Staying Engaged with Chemistry Semester 1 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Semester 1 Review Answers
- 9. Balancing eBooks and Physical Books Chemistry Semester 1 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Semester 1 Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Semester 1 Review Answers
 - Setting Reading Goals Chemistry Semester 1 Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Semester 1 Review Answers
 - Fact-Checking eBook Content of Chemistry Semester 1 Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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