

NCERT Solution for Class 12  
Chemistry  
Chapter 3 - Chemical Kinetics

Intext Exercise

1. For the reaction  $R \rightarrow P$ , the concentration of a reactant changes from 0.03 M to 0.02 M in 25 minutes. Calculate the average rate of reaction using units of time both in minutes and seconds.

Ans: The average rate of the reaction can be calculated by dividing the change in the rate of decreasing the rate of reactant by the time taken. This is given below:

$$\text{Average rate} = - \frac{d[R]}{dt}$$

This can be written as:

$$\text{Average rate} = - \frac{[R]_2 - [R]_1}{t_2 - t_1}$$

$$R_1 = 0.03 \text{ M}$$

$$R_2 = 0.02 \text{ M}$$

$$t_2 - t_1 = 25 \text{ min}$$

Putting the values, we get:

$$\text{Average rate} = - \frac{0.02 - 0.03}{25}$$

$$\text{Average rate} = - \frac{-0.01}{25} = 4 \times 10^{-4} \text{ M min}^{-1}$$

So, the average rate in minutes will be  $4 \times 10^{-4} \text{ M min}^{-1}$

Now, to find the average rate in seconds we have to divide the above answer by 60.

So, the answer will be:

$$\text{Average rate} = - \frac{4 \times 10^{-4}}{60} = 6.66 \times 10^{-6} \text{ M s}^{-1}$$

Therefore, the average rate in seconds will be  $6.66 \times 10^{-6} \text{ M s}^{-1}$

# Chapter 4 Assessment Chemistry Answers Pearson

**S Nieto**



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