

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_ a. ionic bond  
 \_\_\_\_\_ **bond energy** \_\_\_\_\_ b. covalent bond  
 \_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_ a. H—F \_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_ d. H—H  
 \_\_\_\_\_ **ionic** \_\_\_\_\_ b. Na—Cl \_\_\_\_\_ **polar covalent** \_\_\_\_\_ e. H—C  
 \_\_\_\_\_ **polar covalent** \_\_\_\_\_ c. H—O \_\_\_\_\_ **polar covalent** \_\_\_\_\_ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

# Chapter 8 Covalent Bonding Chapter Assessment Answers

**JR Anderson**



**Chapter 8 Covalent Bonding Chapter Assessment Answers:**

## Unveiling the Power of Verbal Beauty: An Mental Sojourn through **Chapter 8 Covalent Bonding Chapter Assessment Answers**

In some sort of inundated with displays and the cacophony of fast interaction, the profound energy and mental resonance of verbal beauty frequently fade into obscurity, eclipsed by the constant assault of sound and distractions. However, set within the lyrical pages of **Chapter 8 Covalent Bonding Chapter Assessment Answers**, a captivating work of fictional brilliance that pulses with raw emotions, lies an memorable journey waiting to be embarked upon. Penned by way of a virtuoso wordsmith, that mesmerizing opus instructions viewers on an emotional odyssey, softly exposing the latent possible and profound affect stuck within the intricate web of language. Within the heart-wrenching expanse of this evocative analysis, we can embark upon an introspective exploration of the book is main styles, dissect their fascinating writing design, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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