

## Chemical Equations and Reactions

### SECTION 2

**SHORT ANSWER** Answer the following questions in the space provided.

1. Match the equation type on the left to its representation on the right.

- |                                  |                                   |
|----------------------------------|-----------------------------------|
| <u>  c  </u> synthesis           | (a) $AX + BY \rightarrow AY + BX$ |
| <u>  d  </u> decomposition       | (b) $A + BX \rightarrow AX + B$   |
| <u>  b  </u> single-displacement | (c) $A + B \rightarrow AX$        |
| <u>  a  </u> double-displacement | (d) $AX \rightarrow A + X$        |

2.   c   In the reaction described by the equation  $2Al(s) + 3Fe(NO_3)_2(aq) \rightarrow 3Fe(s) + 2Al(NO_3)_3(aq)$ , iron has been replaced by  
 (a) nitrate. (c) **aluminum.**  
 (b) water. (d) nitrogen.

3.   a   Of the following chemical equations, the only reaction that is both synthesis and combustion is  
 (a)  **$C(s) + O_2(g) \rightarrow CO_2(g)$ .**  
 (b)  $2C_4H_{10}(l) + 13O_2(g) \rightarrow 8CO_2(g) + 10H_2O(l)$ .  
 (c)  $6CO_2(g) + 6H_2O(g) \rightarrow C_6H_{12}O_6(aq) + 6O_2(g)$ .  
 (d)  $C_6H_{12}O_6(aq) + 6O_2(g) \rightarrow 6CO_2(aq) + 6H_2O(l)$ .

4.   b   Of the following chemical equations, the only reaction that is both combustion and decomposition is  
 (a)  $S(s) + O_2(g) \rightarrow SO_2(g)$ .  
 (b)  **$2C_4H_{10}(l) + 13O_2(g) \rightarrow 8CO_2(g) + 10H_2O(l)$ .**  
 (c)  $2H_2O_2(l) \rightarrow 2H_2O(l) + O_2(g)$ .  
 (d)

5. Identify the products when the following substances decompose:

- Its separate elements** \_\_\_\_\_ a binary compound
- Metal oxide + water** \_\_\_\_\_ most metal hydroxides
- Metal oxide + carbon dioxide** \_\_\_\_\_ a metal carbonate
- Water + sulfur dioxide** \_\_\_\_\_ the acid  $H_2SO_3$

6. The complete combustion of a hydrocarbon in excess oxygen yields the products

# Chemical Equations Reactions Section 2 Answers

**Thomas Griffiths**



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## **Table of Contents Chemical Equations Reactions Section 2 Answers**

1. Understanding the eBook Chemical Equations Reactions Section 2 Answers
  - The Rise of Digital Reading Chemical Equations Reactions Section 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equations Reactions Section 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Equations Reactions Section 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equations Reactions Section 2 Answers
  - Personalized Recommendations
  - Chemical Equations Reactions Section 2 Answers User Reviews and Ratings
  - Chemical Equations Reactions Section 2 Answers and Bestseller Lists

5. Accessing Chemical Equations Reactions Section 2 Answers Free and Paid eBooks
  - Chemical Equations Reactions Section 2 Answers Public Domain eBooks
  - Chemical Equations Reactions Section 2 Answers eBook Subscription Services
  - Chemical Equations Reactions Section 2 Answers Budget-Friendly Options
6. Navigating Chemical Equations Reactions Section 2 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Equations Reactions Section 2 Answers Compatibility with Devices
  - Chemical Equations Reactions Section 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Equations Reactions Section 2 Answers
  - Highlighting and Note-Taking Chemical Equations Reactions Section 2 Answers
  - Interactive Elements Chemical Equations Reactions Section 2 Answers
8. Staying Engaged with Chemical Equations Reactions Section 2 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Equations Reactions Section 2 Answers
9. Balancing eBooks and Physical Books Chemical Equations Reactions Section 2 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Equations Reactions Section 2 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Equations Reactions Section 2 Answers
  - Setting Reading Goals Chemical Equations Reactions Section 2 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Equations Reactions Section 2 Answers
  - Fact-Checking eBook Content of Chemical Equations Reactions Section 2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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