4 Atomic Structure
Reviewsheet

· ·

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are ___1___ from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of ___2___ particles changed this theory. We now know that atoms are made up of electrons, which have a ___3___ charge; ___4___, which have a positive charge; and ___5___, which are neutral. The latter two particles are found in the ___6___ of the atom.

It was ___7___ who discovered the nucleus of the atom. The nucleus has a ___8___ charge and it occupies a very small volume of the atom. In contrast, the negatively charged ___9___ occupy most of the volume of the atom.

The number of ___10___ in the nucleus of the atom is the atomic ___11___ of that element. Because atoms are electrically neutral, the number of protons and ___12___ in an atom are equal. The sum of the ___13___ and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of ___14___ in the nucleus. Atoms that have the same number of protons but different mass numbers are called ___15___.

The ___16___ of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are $^{32}_{16}S$ and $^{34}_{16}S$. An atom of the sulfur–32 isotope contains ___17___ protons and ___18___ neutrons. The sulfur–34 isotope has ___19___ protons and ___20___ neutrons.

1. _different_ 4.1
2. _subatomic_ 4.2
3. _negative_ 4.2
4. _protons_ 4.2
5. _neutrons_ 4.2
6. _nucleus_ 4.3
7. _Rutherford_ 4.3
8. _positive_ 4.3
9. _electrons_ 4.3
10. _protons (elec)_ 4.4
11. _number_ 4.4
12. _electrons_ 4.4
13. _protons (elec)_ 4.5
14. _neutrons_ 4.6
15. _isotopes_ 4.6
16. _atomic mass_ 4.7
17. _16_ 4.5
18. _16_ 4.6
19. _16_ 4.6
20. _18_ 4.5

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