

**Chapter 10 test****Matching**

Match each item with the correct statement below.

- |                            |                                      |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition               |
| b. mole                    | e. standard temperature and pressure |
| c. Avogadro's number       | f. empirical formula                 |

- \_\_\_\_\_ 1. the number of representative particles of a substance present in 1 mole of that substance  
\_\_\_\_\_ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists  
\_\_\_\_\_ 3. the SI unit used to measure amount of substance  
\_\_\_\_\_ 4. 0°C and 1 atm  
\_\_\_\_\_ 5. the percent by mass of each element in a compound  
\_\_\_\_\_ 6. the smallest whole number ratio of the atoms in a compound

**Multiple Choice**

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 7. What SI unit is used to measure the number of representative particles in a substance?  
a. kilogram  
b. ampere  
c. kelvin  
d. mole
- \_\_\_\_\_ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C<sub>3</sub>H<sub>7</sub>O?  
a.  $5 \times (6.02 \times 10^{23})$   
b. 5  
c. 35  
d.  $35 \times (6.02 \times 10^{23})$
- \_\_\_\_\_ 9. All of the following are equal to Avogadro's number EXCEPT \_\_\_\_\_  
a. the number of atoms of bromine in 1 mol Br<sub>2</sub>  
b. the number of atoms of gold in 1 mol Au  
c. the number of molecules of nitrogen in 1 mol N<sub>2</sub>  
d. the number of molecules of carbon monoxide in 1 mol CO
- \_\_\_\_\_ 10. How many moles of tungsten atoms are in  $4.8 \times 10^{23}$  atoms of tungsten?  
a.  $8.0 \times 10^{-2}$  moles  
b.  $8.0 \times 10^1$  moles  
c.  $1.3 \times 10^{-1}$  moles  
d.  $1.3 \times 10^{-2}$  moles
- \_\_\_\_\_ 11. How many atoms are in 0.075 mol of titanium?  
a.  $1.2 \times 10^{23}$   
b.  $2.2 \times 10^{24}$   
c.  $6.4 \times 10^2$   
d.  $4.5 \times 10^{22}$
- \_\_\_\_\_ 12. How many molecules are in 2.10 mol CO<sub>2</sub>?  
a.  $2.53 \times 10^{24}$  molecules  
b.  $3.79 \times 10^{24}$  molecules  
c.  $3.49 \times 10^{24}$  molecules  
d.  $1.26 \times 10^{24}$  molecules

# Chemistry Chapter 10 Test Answers

**David Baud**



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