

CHEMISTRY VIVA QUESTIONS WITH ANSWERS CLASS XII

[1] VOLUMETRIC ANALYSIS

Question.1.What is a standard solution?

Answer. A solution whose strength is known is called a standard solution.

Question.2.What is a normal solution?

Answer. A solution containing one gram-equivalent mass of the solute per litre of the solution is called a normal solution.

Question.3.What is the equivalent mass of KMnO_4 when it acts as oxidizing agent in acidic medium?

Answer. KMnO_4 loses 5 electrons per molecule, when it acts as oxidizing agent in the presence of acids. Therefore, its equivalent mass is one-fifth of its molecular mass.
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Question.4.Is sodium hydroxide a primary standard?

Answer. No.

Question.5.Are 'molality' and 'molarity' same?

Answer. No, molality of a solution is defined as the number of moles of solute present in 1000 grams of the solution whereas molarity tells us about the number of moles of the solute present per litre of the solution.

Question.6.What would be the normality of 0.10M KMnO_4 ?

Answer. It will be $0.1 \times 5 = 0.5 \text{ N}$.

Question.7.What volume of 10M HCl must be diluted with water to get 1L of 1M HCl?

Answer. 0.1L.

Question.8.What is the basicity of H_2SO_4 ?

Answer. 2.

Question.9.What is the relationship between normality (N), molarity (M), molecular mass and equivalent mass?

Answer. Normality \times Eq. Mass = Molarity \times Mol. Mass.

Question.10.Why front door of the balance is closed during weighing?

Answer. Opening the front door causes vibrations in the pan due to operator's breath which leads to inaccurate results.

Chemistry Practical Viva Questions With Answers

Thomas Griffiths



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Table of Contents Chemistry Practical Viva Questions With Answers

1. Understanding the eBook Chemistry Practical Viva Questions With Answers
 - The Rise of Digital Reading Chemistry Practical Viva Questions With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practical Viva Questions With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practical Viva Questions With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practical Viva Questions With Answers
 - Personalized Recommendations
 - Chemistry Practical Viva Questions With Answers User Reviews and Ratings
 - Chemistry Practical Viva Questions With Answers and Bestseller Lists

5. Accessing Chemistry Practical Viva Questions With Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Practical Viva Questions With Answers
 - Interactive Elements Chemistry Practical Viva Questions With Answers
8. Staying Engaged with Chemistry Practical Viva Questions With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practical Viva Questions With Answers
9. Balancing eBooks and Physical Books Chemistry Practical Viva Questions With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practical Viva Questions With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practical Viva Questions With Answers
 - Setting Reading Goals Chemistry Practical Viva Questions With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Practical Viva Questions With Answers
 - Fact-Checking eBook Content of Chemistry Practical Viva Questions With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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