

Equilibrium

- **Equilibrium** is a state where the net force on an object is zero.
 - An object at rest or moving with constant velocity is in equilibrium.
- **Newton's First Law** (Law of Inertia): An object at rest stays at rest, and an object in motion stays in motion with the same speed and direction unless acted upon by an unbalanced force.
 - $F_{net} = 0 \Rightarrow a = 0$
 - $v = \text{constant}$
 - $x = v_0 t + \frac{1}{2} a t^2$
- **Free-body diagrams** show all forces acting on an object.
 - F_g (gravity), F_N (normal), F_f (friction), F_s (spring)
- **Newton's Second Law**: $F_{net} = ma$. If $F_{net} = 0$, then $a = 0$.
 - $F_{net} = 0 \Rightarrow a = 0$
- **Newton's Third Law**: For every action, there is an equal and opposite reaction.
 - $F_{12} = -F_{21}$



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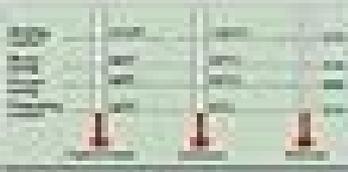
Nuclear Chemistry

- **Radioactivity** is the process by which unstable atomic nuclei lose energy by emitting radiation.
 - α (alpha), β (beta), γ (gamma)
- **Types of Processes**
 - α decay: ${}^A_Z X \rightarrow {}^{A-4}_{Z-2} Y + {}^4_2\text{He}$
 - β decay: ${}^A_Z X \rightarrow {}^A_{Z+1} Y + e^- + \bar{\nu}_e$
 - β^+ decay: ${}^A_Z X \rightarrow {}^A_{Z-1} Y + e^+ + \nu_e$
 - γ decay: ${}^A_Z X^* \rightarrow {}^A_Z X + \gamma$
- **Half-life** is the time it takes for half of a substance to decay.
 - $N = N_0 \left(\frac{1}{2}\right)^{t/T_{1/2}}$
- **Applications**
 - Carbon-14 dating
 - Medical treatments
 - Nuclear power



Measurement & Units

- **SI Units** (International System of Units)
 - Length: meter (m)
 - Mass: kilogram (kg)
 - Time: second (s)
 - Temperature: kelvin (K)
 - Amount of substance: mole (mol)
 - Electric current: ampere (A)
 - Luminous intensity: candela (cd)
- **Prefixes**
 - kilo (k): 10^3
 - centi (c): 10^{-2}
 - milli (m): 10^{-3}
 - micro (μ): 10^{-6}
 - nano (n): 10^{-9}
 - pico (p): 10^{-12}
 - femto (f): 10^{-15}
 - atto (a): 10^{-18}
 - zepto (z): 10^{-21}
 - yocto (y): 10^{-24}
- **Significant Figures**
 - Rules for addition/subtraction
 - Rules for multiplication/division
- **Dimensional Analysis**
 - Check units in calculations



Chemistry Quick Study Academic

TD Snyder



Chemistry Quick Study Academic:

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also considers the interplay of stakeholders consumers and the drug firm with attendant factors including those that are technical legal economic demographic political social ecological and infrastructural Since drug research can be a high risk high payoff industry it is important to researchers to effectively and strategically manage the drug discovery process This book takes a closer look at increasing pre approval costs for new drugs and examines not only why these increases occur but also how they can be overcome to ensure a robust pharmacoeconomic future Written in an engaging manner and including memorable insights this book is aimed at redirecting the drug discovery process to make it more efficient and cost effective in order to achieve the goal of saving countless more lives through science A valuable and compelling resource this is a must read for all students and researchers in academia and the pharmaceutical industry Considers drug discovery in multiple R D venues including big pharma large biotech start up ventures academia and nonprofit research institutes Analyzes the organization of pharmaceutical R D taking into account human resources considerations like recruitment and configuration management of discovery and development processes and the coordination of internal research within and beyond the organization including outsourced work Presents a consistent well connected and logical dialogue that readers will find both comprehensive and approachable

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The Virginia School Journal ,1893 Includes Official department conducted by Superintendent of Public Instruction

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