

# Chemistry – Unit 7 Chemical Reactions

## Lesson 1: Atoms

### Background

Chemistry before 1900 generally focused on the study of elements. The scientific approach to study of the atom is recent.

### Historical View: Classical Model of Matter

The classical theory of matter (particle theory) says that all matter consists of tiny particles which are continuously moving in a continuous space in motion. The frequency of the periodic motion depends on the vibrational energy they have and their relationship to other particles.

### Discovery of Matter

A principle stating that matter has discrete amount of energy:

also in the form of

### Classical Physics

In case of classical systems (classical mechanics) energy is continuous and

has infinite possibilities, while in quantum mechanics the structure of the allowed energy

possibilities are finite and

discrete. The correspondence principle indicates that

number of states of a quantum system is very large in classical limit

and

continuous limit when number of states is very large in quantum limit

A principle or law that describes some property

and

### Procedure

1. The process starts with the reactants (the substances undergoing the chemical reaction) placed in the container. They undergo the reaction to form the product molecules. In all cases, chemical reactions are not subject to direct manipulation (particles involved are tiny and fast-moving).
2. The particles themselves (the reactant molecules) are not visible (cannot be seen) in the process.
3. Chemists use another set of viscosity methods (are visible) in order to study the particles' state in different or complex situations (e.g. direct early chemical reactions).
4. Chemists use particle methods (are visible) in order to study a reaction. It involves measuring the vibrational motion of some atoms in a molecule.

# Chemistry Unit 7 Rearranging Atoms Answers

**Wolfgang Guggemos**



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