

Name - Haanshita Singh ; Class - XII ; Section - A

chapter - Chemical Kinetics - 01.

Advt  
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i) Instantaneous Rate of Reaction.

- (a) It occurs within a short span of time.
- (b) It cannot be calculated for multistep ~~elementary~~ reaction.
- (c) It can be calculated for elementary reaction.

Average Rate of Reaction.

- (a) It occurs during a long interval of time.
- (b) It can be calculated for multistep reaction.
- (c) It cannot be calculated for elementary reaction.

ii) If the unit of rate constant is  $s^{-1}$ , the reaction is of first order.

iii) No, the molecularity can never be zero or a fractional number as it shows the number of reactants taking part in a reaction can never be zero.

iv)



Rate of rxn (M) =  $k[R]^n$  → order of rxn. Reaction involves two reactants.

?  $x = k[R]^n$  — (1)

$8x = k[2R]^n$  — (2)

Dividing eq (2) by (1);

$$\frac{8x}{x} = \frac{k[2R]^n}{k[R]^n}$$

or,  $8 = 2^n \cdot \frac{R^n}{R^n}$

Overall order

= (a+b)

= (1+2) = 3

or,  $8 = 2^n$

or,  $2^3 = 2^n$

$\therefore n = 3$

The order of reaction is 3.

(1)  $x = k[A]^a[B]^b$

(2)  $2x = k[2A]^a[B]^b$

(3)  $8x = k[2A]^a[2B]^b$

(2) ÷ (1) →

$$\frac{2x}{x} = \frac{k[2A]^a[B]^b}{k[A]^a[B]^b}$$

$\sim 2^1 = 2^a \implies a = 1$

(3) ÷ (2) →

$2^b = 4 = 2^2 \implies b = 2$

(v) The rate constant for the first order reaction is  $60 s^{-1}$ .

let M be the molar mass.  
a be the initial concentration.

$k = 60/s$

final concentration =  $\frac{M}{16} \times a$

# Chemsheets Kinetics 1 Answers

**Joacim Rocklöv**



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Introduction to Chemical Kinetics Gordon Skinner, 2012-12-02 Introduction to Chemical Kinetics is a compilation of lecture notes of the author about principles concepts and theories in chemical kinetics The book tackles the nature of chemical kinetics reaction rates and order and thermodynamic consistency of rate laws The effects of temperature on kinetics prediction of reaction rates gas phase reactions and controlled reactions are also discussed The text also explains the reactions catalyzed by enzymes reactions in solids and heterogenous systems oxidation of metals catalysis of reactions by solids and methods for different reaction rates The monograph is recommended as a textbook for undergraduate students in chemistry who are currently taking up kinetics as it is an easily understood and concise book that can also be used as reference

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