

STUDENT BOOK ANSWERS
To come
Chapter 9: Gas laws
Worked example 9.1

- 1 $n = \frac{V}{V_m} = \frac{2.33}{22.71} = 0.1026 \text{ mol}$
 $m = n \times MM = 0.1026 \times (12.01 + (4 \times 1.008)) = 1.646 = 1.65 \text{ g (3sf)}$
- 2 $n = \frac{m}{MM} = \frac{23.0}{14.01 + (3 \times 1.008)} = 1.350 \text{ mol}$
 $V = n \times V_m = 1.350 \times 24.79 = 33.47 = 33.5 \text{ L (3sf)}$

Worked example 9.2

- 1 $2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$
 $n(\text{Na}) = \frac{m}{MM} = \frac{1.87}{22.99 + 16.00 + 1.008} = 0.04675 \text{ mol}$
 $n(\text{H}_2) = \frac{1}{2} \times n(\text{Na}) = 0.5 \times 0.04675 = 0.02338 \text{ mol}$
 $V = n \times V_m = 0.02338 \times 24.79 = 0.5795 = 0.580 \text{ L (3sf)}$
- 2 $n(\text{HCl}) = \frac{V}{V_m} = \frac{13.4}{22.71} = 0.5900 \text{ mol}$
 $2\text{KCl(aq)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow 2\text{HCl(g)} + \text{K}_2\text{SO}_4\text{(aq)}$
 $n(\text{KCl}) = n(\text{HCl}) = 0.5900 \text{ mol}$
 $m(\text{KCl}) = n \times MM = 0.5900 \times (39.10 + 35.45) = 43.98 = 44.0 \text{ g (3sf)}$

Check your understanding 9.1/9.2

- 1 The volume of a gas changes with temperature and pressure, both of which compress or expand the gas particles.
- 2 Use the formula $n = \frac{V}{V_m}$ to convert the volume of gas to a number of moles. V_m depends upon the temperature of the gas. Use $m = n \times MM$ to convert the number of moles (n) to a mass (m) by multiplying by the molecular mass (MM).
- 3 **a** **i** $2 \times 250 = 500 \text{ mL}$
ii 250 mL
b **i** 500 mL
ii Nothing, as the law of combining volumes only applies to gases.

Chapter 9 Notes Cengage Learning

David Deming



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