

Chapter 6

Equilibrium

Question 6.47:

It has been found that the pH of a 0.01M solution of an organic acid is 4.15. Calculate the concentration of the anion, the ionization constant of the acid and its pK_a .

Answer 6.47:

Let the organic acid be HA.



Concentration of HA = 0.01 M pH = 4.15

$$-\log[H^+] = 4.15$$

$$[H^+] = 7.08 \times 10^{-5}$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

Now,

$$[H^+] = [A^-] = 7.08 \times 10^{-5}$$

$$[HA] = 0.01$$

Then,

$$K_a = \frac{(7.08 \times 10^{-5})(7.08 \times 10^{-5})}{0.01} = 5.01 \times 10^{-7}$$

$$pK_a = -\log K_a = -\log(5.01 \times 10^{-7}) = 6.3001$$

Question 6.48:

Assuming complete dissociation, calculate the pH of the following solutions:

(i) 0.003 M HCl

(ii) 0.005 M NaOH

(iii) 0.002 M HBr

(iv) 0.002 M KOH

Answer 6.48:

(i) 0.003M HCl:



Since HCl is completely ionized,

$$[H_3O^+] = [HCl] = 0.003$$

Now,

$$pH = \log[H_3O^+] = -\log(.003) = 2.52$$

Hence, the pH of the solution is 2.52.

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Table of Contents Chemistry Chapter 6 Solutions Manual

1. Understanding the eBook Chemistry Chapter 6 Solutions Manual
 - The Rise of Digital Reading Chemistry Chapter 6 Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 6 Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 6 Solutions Manual
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chemistry Chapter 6 Solutions Manual
 - Personalized Recommendations
 - Chemistry Chapter 6 Solutions Manual User Reviews and Ratings
 - Chemistry Chapter 6 Solutions Manual and Bestseller Lists
5. Accessing Chemistry Chapter 6 Solutions Manual Free and Paid eBooks
 - Chemistry Chapter 6 Solutions Manual Public Domain eBooks
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 - Chemistry Chapter 6 Solutions Manual Budget-Friendly Options
6. Navigating Chemistry Chapter 6 Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 6 Solutions Manual Compatibility with Devices
 - Chemistry Chapter 6 Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 6 Solutions Manual
 - Highlighting and Note-Taking Chemistry Chapter 6 Solutions Manual
 - Interactive Elements Chemistry Chapter 6 Solutions Manual
8. Staying Engaged with Chemistry Chapter 6 Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 6 Solutions Manual
9. Balancing eBooks and Physical Books Chemistry Chapter 6 Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 6 Solutions Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 6 Solutions Manual
 - Setting Reading Goals Chemistry Chapter 6 Solutions Manual
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Chapter 6 Solutions Manual
 - Fact-Checking eBook Content of Chemistry Chapter 6 Solutions Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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