

Preparation of a standard solution

Aim: To prepare 2.0 mol.dm^{-3} standard sodium chloride solution using a 250 cm^3 volumetric flask.

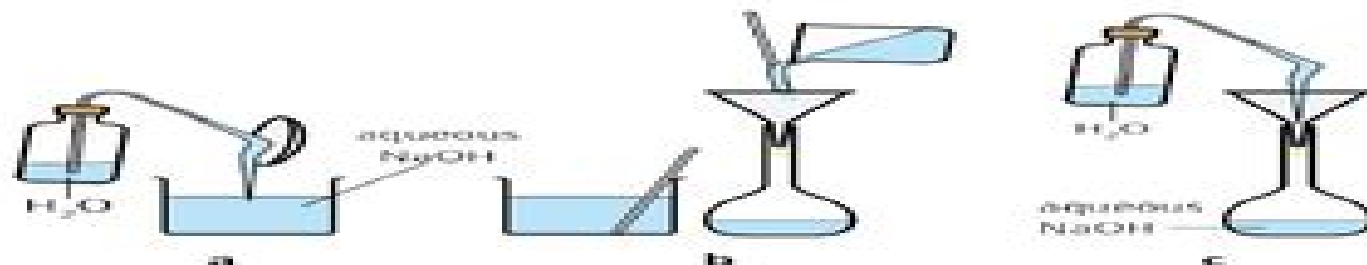
Tools 1: Experimental techniques – Applying techniques.

Available Apparatus: digital balance, volumetric flask, pipettes, and pipette bulbs, spatula, watch glass, beaker, funnel, stirring rod, and wash bottle.

Reagents: Sodium chloride, food coloring.

Preparation of sodium hydroxide solution.

1. Calculate the mass (in grams) of solid sodium hydroxide required to prepare 2.0 mol.dm^{-3} solution in the assigned volumetric flask.
2. Place a watch glass on the analytical balance and press the 'Tare' button.
3. Accurately weigh approximately the amount of sodium chloride, you have calculated you need, recording the exact amount used.
4. Add approximately 50 cm^3 distilled water to the beaker and stir with a glass rod until the solid has dissolved.
5. Using a funnel, carefully transfer your solution to the volumetric flask given.
6. Add approximately 10 cm^3 distilled water to the beaker, rinse it around the beaker and add to the volumetric flask. Repeat with a further 10 cm^3 portion.
7. Carefully add distilled water to the volumetric flask, running it around the edge of the funnel to wash any remaining NaCl in, and also running it over the stirring rod to wash this in.
8. When the level of solution is within 1 cm of the graduation on the flask, add distilled water drop-wise using a pipette until the bottom of the meniscus is in line with the graduation mark.
9. Firmly place the lid on the volumetric flask, and invert it several times, holding the lid in place. You will know your solution is fully mixed when inverting it no longer makes it look 'swirly'.
10. Calculate the exact concentration (4 s.f.) of the solution, using the mass measured in Step 2 and label the flask with the solution, the concentration, the date of preparation and your name.
11. The solution should be stored in a cool dark place until it is to be used.



A. Calculate the mass of sodium chloride required to prepare a solution 2.0 mol.dm^{-3} solution.

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Preparation Of Solutions Lab Report

Theodore Weicker



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Table of Contents Preparation Of Solutions Lab Report

1. Understanding the eBook Preparation Of Solutions Lab Report
 - The Rise of Digital Reading Preparation Of Solutions Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Preparation Of Solutions Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Preparation Of Solutions Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Preparation Of Solutions Lab Report
 - Personalized Recommendations
 - Preparation Of Solutions Lab Report User Reviews and Ratings
 - Preparation Of Solutions Lab Report and Bestseller Lists
5. Accessing Preparation Of Solutions Lab Report Free and Paid eBooks
 - Preparation Of Solutions Lab Report Public Domain eBooks
 - Preparation Of Solutions Lab Report eBook Subscription Services
 - Preparation Of Solutions Lab Report Budget-Friendly Options

6. Navigating Preparation Of Solutions Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Preparation Of Solutions Lab Report Compatibility with Devices
 - Preparation Of Solutions Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Preparation Of Solutions Lab Report
 - Highlighting and Note-Taking Preparation Of Solutions Lab Report
 - Interactive Elements Preparation Of Solutions Lab Report
8. Staying Engaged with Preparation Of Solutions Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Preparation Of Solutions Lab Report
9. Balancing eBooks and Physical Books Preparation Of Solutions Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Preparation Of Solutions Lab Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Preparation Of Solutions Lab Report
 - Setting Reading Goals Preparation Of Solutions Lab Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Preparation Of Solutions Lab Report
 - Fact-Checking eBook Content of Preparation Of Solutions Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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








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