

26. The number of significant figures in the measured value 0.0204 is  
 (1) Five (2) Three (3) Four (4) Two

**Sol. Answer (2)**

The non-zero digits after the decimal places are significant.

27. The number of significant figures in the measured value 26000 is  
 (1) Five (2) Two (3) Three (4) Infinite

**Sol. Answer (2)**

The trailing zeros are not significant.  
 So, only two digits are significant.

28. The number of significant zeroes present in the measured value 0.020040, is  
 (1) Five (2) Two (3) One (4) Three

**Sol. Answer (4)**

Zeros appearing between and after non-zero numbers are significant.

0.020040

29. The number of significant figures in the measured value 4.700 m is the same as that in the value  
 (1) 4700 m (2) 0.047 m (3) 4070 m (4) 470.0 m

**Sol. Answer (4)**

4.700  $\Rightarrow$  Four significant figures.

Also, 470.0 m  $\Rightarrow$  Four significant figures.

30. If a calculated value 2.7465 g contains only three significant figures, the two insignificant digits in it are  
 (1) 2 and 7 (2) 7 and 4 (3) 6 and 5 (4) 4 and 6

**Sol. Answer (3)**

2.7465 g  $\Rightarrow$  Last two digits are most insignificant.

31. An object of mass 4.237 g occupies a volume 1.72 cm<sup>3</sup>. The density of the object to appropriate significant figures is

- (1) 2.46 g cm<sup>-3</sup> (2) 2.463 g cm<sup>-3</sup> (3) 2.5 g cm<sup>-3</sup> (4) 2.50 g cm<sup>-3</sup>

**Sol. Answer (1)**

$$m = 4.237 \text{ g}$$

$$V = 1.72 \text{ cm}^3$$

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}} = \frac{4.237 \text{ g}}{1.72 \text{ cm}^3}$$

$$\Rightarrow \boxed{d = 2.46 \text{ g cm}^{-3}}$$

32. Round off the value 2.845 to three significant figures.

- (1) 2.85 (2) 2.84 (3) 2.80 (4) 2.83

**Sol. Answer (2)**

$$2.845 \Rightarrow 2.84$$

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