

Inorganic Chemistry

Transition Metals

Introduction

Transition metals are elements in the d-block of the periodic table, characterized by their ability to form complex ions and exhibit variable oxidation states. They are found in groups 3 to 10 of the periodic table.

The d-orbitals in transition metals are partially filled, leading to their characteristic properties such as color, catalytic activity, and the ability to form complex ions. The crystal field theory (CFT) explains the splitting of d-orbitals in a complex, which is responsible for the color of many transition metal complexes.

Transition metals are also known for their high melting and boiling points, and their ability to form alloys. The high strength of these metals is due to the strong metallic bonding between the d-orbitals.

Crystal Field Theory

Crystal field theory (CFT) is a model that describes the interaction between the d-orbitals of a transition metal ion and the ligands in a complex. The ligands are treated as point charges, and their electric field causes the degenerate d-orbitals to split into different energy levels. The energy difference between these levels is called the crystal field splitting energy (Δ_o).

Coordination Chemistry

Coordination Compounds

Coordination compounds are formed by the central metal ion and the surrounding ligands. The central metal ion is typically a transition metal ion, and the ligands are molecules or ions that donate lone pairs of electrons to the metal ion.

Coordination Number

The coordination number of a central metal ion is the number of ligands that are bonded to it. It is determined by the size of the metal ion, the size of the ligands, and the geometry of the complex. Common coordination numbers are 2, 4, and 6.

The geometry of a coordination complex is determined by the coordination number and the nature of the ligands. Common geometries include octahedral, tetrahedral, and square planar. The geometry affects the properties of the complex, such as its color and magnetic behavior.

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Organometallic Chemistry

Organometallic Compounds

Organometallic compounds are compounds that contain a metal atom bonded to one or more carbon atoms. They are found in a wide range of applications, including catalysis, materials science, and medicine.

Coordination Chemistry

Organometallic compounds often form complex ions, and their properties are determined by the nature of the metal atom and the ligands. The coordination number and geometry of the complex are important factors in determining the properties of the compound.

Organometallic compounds are used in a wide range of applications, including catalysis, materials science, and medicine. They are also used in the synthesis of many organic compounds.

Crystal Field Theory

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